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Measurement of gas-phase ammonia and amines in air by collection onto an ion exchange resin and analysis by ion chromatography

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Abstract

Ammonia and amines are common trace gases in the atmosphere and have a variety of both biogenic and anthropogenic sources, with a major contribution coming from agricultural sites. In addition to their malodorous nature, both ammonia and amines have been shown to enhance particle formation from acids such as nitric, sulfuric and methanesulfonic acids, which has implications for visibility, human health and climate. A key component of quantifying the effects of these species on particle formation is accurate gas-phase measurements in both laboratory and field studies. However, these species are notoriously difficult to measure as they are readily taken up on surfaces, including onto glass surfaces from aqueous solution as established in the present studies. We describe here a novel technique for measuring gas-phase ammonia and amines that involves uptake onto a weak cation exchange resin followed by extraction and analysis using ion chromatography. Two variants, one for ppb concentrations in air and the second with lower (ppt) detection limits, are described. The latter involves the use of a custom-designed high-pressure cartridge to hold the resin for in-line extraction. These methods avoid the use of sampling lines, which can lead to significant inlet losses of these compounds. They also have the advantages of being relatively simple and inexpensive. The applicability of this technique to ambient air is demonstrated in measurements made near a cattle farm in Chino, CA.

20 1 Introduction

Atmospheric aerosol particles are known to reduce visibility and adversely affect human health. They also impact the climate as they are able to scatter and absorb solar radiation and serve as cloud and ice condensation nuclei (Finlayson-Pitts and Pitts, 2000; Seinfeld and Pandis, 2006). Ammonia and amines are routinely detected in the particle phase and have been identified as important contributors to new particle formation and growth (Angelino et al., 2001; Berndt et al., 2010; Bzdek and Johnston,

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2010; Bzdek et al., 2011; Chen et al., 2012; Creamean et al., 2011; Dawson et al., 2012; Loukonen et al., 2010; Muller et al., 2009; Ruiz-Jiménez et al., 2012; Smith et al., 2010; VandenBoer et al., 2011). Accurate measurement of gas-phase ammonia and amines both in the atmosphere and in laboratory experiments is a key component 5 of understanding and quantifying their role in particle chemistry and physics.

Atmospheric ammonia and amines have a wide variety of sources, both biogenic and anthropogenic (Ge et al., 2011a, b). Industrial and agricultural practices involving animals, e.g., cattle feed lots or swine facilities (Anderson et al., 2003; Hiranuma et al., 2010; Kuhn et al., 2011; Mosier et al., 1973; Ni et al., 2012; Schade and Crutzen, 1995), 10 are significant sources of these species in the atmosphere. Agricultural emissions of ammonia and short-chain aliphatic amines, in particular, account for a large fraction of the global flux of these species into the atmosphere (Ge et al., 2011a, b; Schade and Crutzen, 1995). Other sources include ocean biota (Ge et al., 2011a; Gibb et al., 1999), biomass burning (Ge et al., 2011a; Lobert et al., 1990), and release from carbon 15 capture and storage devices that use amines to trap CO_2 , which could be a more important source of atmospheric amines and ammonia as the technology becomes more widely adopted (Borduas et al., 2013; Ge et al., 2011a; Nielsen et al., 2012; Rochelle, 2009; Schreiber et al., 2009).

Importantly, it has been shown that gas-phase ammonia and amines significantly 20 enhance particle formation from common atmospheric acids, such as sulfuric, nitric and methanesulfonic acids (Almeida et al., 2013; Angelino et al., 2001; Berndt et al., 2010; Chen et al., 2012; Dawson et al., 2012; Loukonen et al., 2010; Smith et al., 2010; Yu et al., 2012), and contribute to growth of atmospheric nano-particles (Barsanti 25 et al., 2009; Bzdek et al., 2011; Smith et al., 2010; Wang et al., 2010). Short-chain alkyl amines have been shown to displace ammonia in particles (Bzdek et al., 2010, 2011; Chan and Chan, 2012; Liu et al., 2012; Lloyd et al., 2009), which enhances their importance in particle formation and growth. Although amines are short-lived in the atmosphere due to oxidation by common atmospheric oxidants such as OH and O_3 (Nielsen et al., 2012; Tang et al., 2013), amines and the precursors to sulfuric and

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methanesulfonic acids (Bates et al., 1992; Ge et al., 2011a; Ni et al., 2012) are in sufficiently close temporal proximity in the atmosphere that they are important contributors to particle formation.

Reliable data on the sources, sinks, and ambient concentrations of gas-phase amines, therefore, are crucial to predicting new particle formation in the atmosphere. However, gas-phase amines are notoriously difficult to measure and typical concentrations in the atmosphere are of the order of a few ppb or less (Ge et al., 2011a). Several techniques for measuring gas-phase ammonia and amines have been reported in the literature. On-line mass spectrometric (MS) techniques include ambient pressure proton transfer MS (Hanson et al., 2011), chemical ionization MS (Yu and Lee, 2012), and proton transfer reaction MS (PTR-MS) (Borduas et al., 2013; Feilberg et al., 2010; Kuhn et al., 2011; Liu et al., 2011; Tanimoto et al., 2007). Off-line techniques typically involve collection of a gas-phase sample onto a substrate (e.g., activated charcoal or an acid-impregnated glass fiber filter) (Fournier et al., 2008; Fuselli et al., 1982), into an acidic solution (Akyüz, 2008; Gronberg et al., 1992; Schade and Crutzen, 1995), or onto a whetted glass frit (Huang et al., 2009). Samples are then extracted and analyzed using gas or liquid chromatography, sometimes with a derivatization step included (Akyüz, 2008; Fuselli et al., 1982; Gronberg et al., 1992; Hiranuma et al., 2010; Huang et al., 2009; Nishikawa and Kuwata, 1984; Santagati et al., 2002; Schade and Crutzen, 1995). Ion chromatography (IC) has also proven to be useful for both gas- and particle-phase ammonia and amines (Gibb et al., 1999; Orsini et al., 2003; Praplan et al., 2012; VandenBoer et al., 2011). Formation of an indophenol complex which is measured spectrometrically has been developed for NH_3 (Hiranuma et al., 2010), as have various techniques involving the formation of 1-sulfonatoisoindole followed by fluorescence measurement (Toda et al., 2010; Zhang et al., 1989).

An important limitation to many existing techniques for measuring ammonia and amines is deposition of the gas-phase analyte onto instrument surfaces prior to measurement, which varies with the compound (Hansen et al., 2013). Also, it has recently been shown that amines are irreversibly taken up onto surfaces that have been exposed

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to a gas-phase acid, forming a non-volatile salt (Nishino et al., 2013). As a variety of acids and acid precursors are present in the atmosphere, this loss may have a significant effect on measurement efficiency for instrumentation where the gas-phase sample is in contact with surfaces such as tubing prior to measurement, even when these surfaces are heated. In addition, amines in aqueous solution are shown here to be subject to uptake on glass, with implications for sampling and measurement.

This work demonstrates the use of a weak cation exchange (WCE) resin as a substrate for efficient collection of gas-phase ammonia and amines at atmospherically relevant concentrations, followed by analysis by IC. Two approaches were developed.

10 The first is applicable to higher (ppb) concentrations while the second, for which a custom high-pressure resin holder cartridge was designed for in-line extraction on an IC system, has detection limits in the tens of ppt range. These methods were developed to minimize the sampling losses reported previously, whilst also being capable of measuring ammonia and amines at the ppt level in air. Application to both laboratory and field measurements is demonstrated.

2 Experimental

2.1 Liquid-phase standards

Standard solutions for ammonia, methylamine (MA), dimethylamine (DMA), and trimethylamine (TMA) were prepared from their chloride salts in 0.1 M oxalic acid (Fluka). These include NH_4Cl (Sigma, 99.5 %), $\text{CH}_3\text{NH}_3\text{Cl}$ (Aldrich, 98.0 %), $(\text{CH}_3)_2\text{NH}_2\text{Cl}$ (Aldrich, 99.0 %), and $(\text{CH}_3)_3\text{NHCi}$ (Aldrich, 98.0 %).

In the course of developing this method, there was some indication that amines and/or aminium ions in aqueous solution were being taken up in the walls of glass containers. To test whether, and to what extent, this was occurring, a standard solution containing between 10 and 30 ng mL^{-1} of the ammonium and aminium species in nanopure water was placed in three 20 mL glass scintillation vials half filled with clean,

dry glass beads and allowed to sit for 60 min. The original standard solution and those from the glass vials were then analyzed by IC. The peaks in the samples from the glass vials corresponding to ammonia and the amines were reduced, on average, by 13–23 % compared to the original standard solution, indicating that amines are taken up by glass surfaces under neutral conditions. However, it should be noted that if the standard solution was acidic, no uptake on glass was observed. To avoid any potential wall loss, no glass was used in the preparation or storage of standards and samples used in this study.

2.2 Gas-phase standards

10 Mixtures of ammonia (Airgas; 0.812 ppm in N₂), MA (Airgas; 10 ppm in N₂), DMA (Airgas; 1.0 ppm in N₂), and TMA (Airgas; 1.0 ppm in N₂) in nitrogen were used to test the collection efficiency of the cation exchange resin (stated concentrations were those provided by the manufacturer but as discussed below, have considerable uncertainties associated with them). Gas-phase ammonia and amines from the gas cylinders were 15 diluted with clean, dry air from an FTIR purge air generator (Parker-Balston; Model 75–62) for a total flow of 4.0 L min⁻¹ and analyte concentrations of approximately 2–1000 ppb as shown in Fig. 1. Gas cylinder and purge-air flows were maintained using mass flow controllers (Alicat).

2.3 Cartridge preparation and analysis for higher (ppb) concentrations

20 Sampling cartridges were prepared by filling 2.5 mL non-fluorous polypropylene cartridges (Supelco; model 57602-U) with WCE resin (Resintech, model WACG) between two polyethylene frits (Supelco) as shown in Fig. 1 (inset). WCE resin consists of acrylic/divinylbenzene beads terminated with carboxylic acid groups. The design of the cartridges minimizes the surfaces in contact with the sample prior to adsorption on the WCE resin, and those that are exposed are subsequently extracted with the resin. 25 These cartridges were used to sample gas-phase standards in the ppb–ppm range in

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air to characterize the collection and extraction efficiency of WCE resin, and will be referred to as “high-concentration cartridges”.

Samples were collected for 20 min at $1\text{--}2\text{ L min}^{-1}$ maintained using a mass flow controller (Alicat). Two cartridges in series (Fig. 1, hereafter referred as “primary” and “backup” cartridges) were used in all experiments to determine collection efficiency. Cartridges were extracted and regenerated by flushing five times with 10.0 mL 0.1 M oxalic acid (Fluka) to remove the collected ammonia and amines and return the resin to its protonated (-R-COOH) form. The first 10.0 mL extract was used as the sample. For the primary cartridge, the first two 10.0 mL extractions were analyzed to determine extraction efficiency. The cartridge was flushed another 3 times with 10.0 mL 0.1 M oxalic acid and the final rinse was used as a blank for the subsequent sample. Samples were stored in 11 mL polypropylene vials (Metrohm, KITIC0008) prior to measurement.

Samples were analyzed by IC (Metrohm, model 850) with a Metrosep C4 – 250/4.0 cation column and equipped with a conductivity detector. The IC eluent was 0.00375 M oxalic acid and the flow rate was 0.9 mL min^{-1} . The IC column temperature was maintained at $30\text{ }^\circ\text{C}$. The sample loop was 20 μL , and the total elution time was 24 min.

2.4 Cartridge preparation and analysis for lower (ppt) concentrations by in-line extraction and analysis

For ambient sampling, modified cartridges (Fig. 2) that could be used under the high-pressure conditions of the IC were designed for gas-phase collection and in-line extraction (see below). It should be noted that “in-line” here refers to the method of extraction on the IC column and does not indicate that this is an on-line measurement technique. These cartridges were prepared using a PEEK analytical guard cartridge holder (Hamilton; model 79477) designed for use on high-pressure liquid chromatography systems and a custom-built stainless-steel insert containing WCE resin (Resintech, model WACG) between two polyethylene frits (Supelco). These are referred to as “low-concentration cartridges” in the subsequent discussion. As for the “high-concentration” cartridge, this design minimizes the amount of surface area that gas-phase samples

are in contact with prior to adsorption on the WCE resin to $\sim 1\text{ cm}^2$ of stainless steel and one of the polyethylene frits; however, adsorbates on both the frit and stainless steel are extracted along with those on the resin.

Prior to sampling, the low-concentration cartridges were flushed three times with 10.0 mL 0.1 M oxalic acid followed by clean, dry air for 20 min at $150\text{ cm}^3\text{ min}^{-1}$ to remove residual water from the last rinse. Gas-phase samples were pumped through the cartridge at $150\text{ cm}^3\text{ min}^{-1}$ for 45–50 min in the direction indicated in Fig. 2. After sampling, the cartridge was filled with 60–80 μL 0.00375 M oxalic acid (IC eluent) in the same direction as the gas-phase sample (Fig. 2), to avoid injecting air into the IC system.

Extraction and analysis were performed in-line on the IC by using two injectors in series, as shown in Fig. 3. This procedure eliminates the separate extraction step and allows the entire collected sample to be injected onto the IC column, as opposed to extracting the cartridge with 10 mL of 0.1 M oxalic acid and then analyzing a 20 μL portion of the extract on the IC. Having the entire collected sample injected onto the IC column lowers the detection limit to a range suitable for atmospheric concentrations (Table 1). The first injector sample loop was loaded with 20 μL 0.1 M oxalic acid and the second injector was fitted with the low-concentration cartridge in place of a sample loop, oriented so the IC eluent flow will be in the direction indicated in Fig. 2. All other IC conditions were as described in Sect. 2.3.

At the beginning of the run, the sample loop containing the acid and the low-concentration cartridge were simultaneously injected. This allows the concentrated oxalic acid plug to extract the cartridge and push the analyte onto the column. After 0.25 min, the cartridge injector was returned to fill mode. Three to five sequential extractions of the cartridge were performed for each sample, depending on the measured ammonia and/or amine concentrations.

A series of experiments was performed to determine if breakthrough occurs in the low-concentration cartridges under conditions of high ammonia concentration as is often seen in the field samples (Ge et al., 2011a, b). Three low-concentration cartridges

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were prepared as described above, and one was kept as a blank. Gas-phase ammonia in N₂ from a gas cylinder (Airgas; 0.812 ppm in N₂) was then flowed through the remaining two cartridges in series for 50 min at 150 cm³ min⁻¹. Three sets of samples were taken. After background subtraction, the measured NH₃ concentration on the backup cartridge was less than 7 % of the total measured NH₃ (primary + backup) in all three cases, and was 4 % of the total measured NH₃ on average. These results suggest minimal breakthrough occurs, even with high ammonia concentrations.

2.5 Field measurements in an agricultural area

Several field measurements using the low-concentration cartridges were performed in Chino, CA, USA. A 40 L steel chamber under vacuum was used as the pump. It was evacuated, fitted with a battery-powered mass flow controller (Alicat) and used to maintain sample flow through the cartridges. This allowed sampling to be performed away from a power source without the need for a generator, which could have introduced exhaust-related artifacts. Samples were taken approximately 50 m away from cattle pens between 4 a.m. and 6 a.m. (before sunrise) between 28 August and 12 September 2013. On each day, one cartridge was prepared as described in Sect. 2.4 and kept as a blank. These blanks were used for background subtraction of the sample chromatograms.

3 Results and discussion

A typical chromatogram for the liquid standards is shown in Fig. 4. Peaks corresponding to NH₄⁺, MA-H⁺, DMA-H⁺, TMA-H⁺ and a small amount of Na⁺ are present. Table 1 summarizes retention times and liquid-phase detection limits for ammonia, MA, DMA and TMA, calculated from the average of the signal corresponding to 3/5 peak-to-peak noise from 10 typical cartridge measurements. For the high-concentration cartridges, gas-phase detection limits were calculated for 20 min samples at 1.0 L min⁻¹ sample

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flow followed by extraction in 10 mL 0.1 M oxalic acid and injection of 20 μL of this extract in the IC. For the low-concentration cartridges, gas-phase detection limits were calculated for 60 min samples at $150 \text{ cm}^3 \text{ min}^{-1}$ sample flow followed by in-line extraction on the IC. Detection limits for the high- and low-pressure cartridges were in the low ppb and ppt range, respectively (see Table 1). It should be noted that with the current design of the low-concentration cartridges, $150 \text{ cm}^3 \text{ min}^{-1}$ is the maximum sample flow possible, however redesigning the cartridge to allow higher sampling flow would further lower the detection limits for this method.

3.1 Gas-phase standards using the high-concentration cartridges with off-line extraction

Oxalic acid is not retained by the cation column used in the IC and elutes at ~ 2.5 min. The high concentration (0.1 M) of oxalic acid in the cartridge extracts compared to that of the IC eluent (0.00375 M) results in a characteristic negative broad signal initially as can be seen in a typical cartridge blank (Fig. 5a). For this reason, blanks are subtracted from cartridge samples before the peaks are integrated. A typical background-subtracted chromatogram for a DMA sample is shown in Fig. 5b. Results from the gas-phase standard measurements are presented in Fig. 6 and show measured ammonia and amine concentrations for the first and second extract of the primary cartridge, the first extract of the backup cartridge, as well as the total measured concentration (first and second extract of the primary cartridge plus the first extract of the backup cartridge).

WCE resin was originally designed to remove alkaline components from liquid solutions by reaction with the surface carboxylic acid groups (Kunin and Barry, 1949). To the best of our knowledge, its ability to take up gas-phase species has not been reported. For the three amines, the measured concentration from the backup cartridge was less than 5 % of that of the primary cartridge (Fig. 6). This small amount of breakthrough indicates that WCE resin efficiently takes up gas-phase amines even at the relatively high flow rate of 1.0 L min^{-1} . For ammonia, this value is slightly higher (< 10 % of the

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negligible amounts measured in the backup cartridge and second extract of the primary cartridge indicate efficient measurement for this technique. This comparison of measured to nominal concentrations also provides a cautionary note in terms of using commercially supplied amine or ammonia gas mixtures as calibration standards.

5 All the samples for the gas-phase standard measurements (both primary and backup) were collected on four high-concentration cartridges. These cartridges showed no noticeable degradation in collection or extraction efficiency, even after hundreds of extractions without replacing the WCE resin.

3.2 Results for field measurements using low-concentration cartridges

10 A typical chromatogram from an air sample taken in Chino, CA, on 28 August 2013 is shown in Fig. 7. On each of the three days of sampling (28 August, 04 and 12 September 2013), two 45–50 min samples were taken. The results from all field measurements are presented in Table 2. Also included in Table 2 are the temperature, relative humidity and weather conditions for each sample as reported by NOAA for the Chino Airport, 15 which is < 1 mile away from the sampling site (NOAA, 2014).

In all samples, peaks corresponding to NH_3 and TMA were observed, and in several samples a peak for MA and/or a peak at ~ 14 min were present. In addition to the standards described in Sect. 2.1, those for isopropylamine, ethylamine, diethylamine, butylamine and aniline were obtained and analyzed by IC, however their retention times 20 did not correspond to the peak at ~ 14 min, which remains unidentified. Also, diethylamine has been reported to coelute with TMA in some Dionex IC columns (Murphy et al., 2007; VandenBoer et al., 2011, 2012). However, using the Metrohm column and IC conditions described in Sect. 2.3, these two species were sufficiently well resolved to identify.

25 Each sample cartridge was extracted in-line and analyzed sequentially by IC five times. The TMA peaks for the five extracts from a sample take on 28 August 2013 are shown in Fig. 8. The trend in integrated peak areas with extraction for TMA (shown in parentheses in Fig. 8) indicates that five extractions are necessary to measure > 97 %

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of the collected species. Results for ammonia show the same trend. While ammonia and TMA peaks were usually still present in the fifth extract, they represented $3 \pm 2\%$ (2s) for NH_3 and $1 \pm 3\%$ (2s) for TMA of the total over five extractions. However, the need for five extractions (~ 2.5 h IC run time) is a limitation of this method over existing on-line techniques. Optimization of this method (e.g. modifying cartridge dimensions, extraction solvent, IC parameters, etc.) may be able to reduce the number of required extractions, thereby improving the efficiency of this technique.

The first measurement on 12 September 2013 showed noticeably lower ammonia and TMA concentrations compared to the previous sampling periods. Several factors may contribute to this difference. The temperature was lower and the relative humidity higher on 12 September compared to the two previous sampling days. Also, mist was reported by the Chino Airport weather station on 12 September that had completely cleared up sometime between 5:38 a.m. and 5:53 a.m. Wet deposition onto mist droplets could account for the lower NH_3 and TMA concentrations on this day and also explain the increase in concentration between the first and second sample as the mist evaporated.

The results of these field measurements are consistent with the range of published data on ammonia and amine concentrations in agricultural areas. Concentrations of ammonia and TMA near cattle feedlots and enclosures in the range of 0.7–34 ppm NH_3 (Hiranuma et al., 2010; Huang et al., 2009; Trabue et al., 2011) and 0–400 ppt TMA (Fujii and Kitai, 1987; Kuwata et al., 1983; Trabue et al., 2011) have been reported. Inside cattle enclosures, TMA concentrations up to 0.6–7.6 ppb have been measured (Fujii and Kitai, 1987; Kallinger and Niessner, 1999; Kuhn et al., 2011). The ratio of TMA to NH_3 in this study, $(4\text{--}9) \times 10^{-3}$, is similar to that reported for indoor cattle enclosures as well as emissions from hay and silage (Kuhn et al., 2011).

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4 Conclusions

This technique involving weak cation exchange resin as a substrate for collection of gas-phase ammonia and amines offers an accurate, reproducible, and inexpensive means of measurement at atmospherically relevant concentrations that is useful for 5 both laboratory and field studies. It minimizes losses on inlets and sampling lines, and avoids uptake of aqueous amines onto glass surfaces. In addition, it is simple and relatively easy to implement, and uses commonly available instrumentation. The custom-designed high-pressure cartridge used as a carrier for the resin combined with a technique for in-line extraction of the compounds and analysis by ion chromatography 10 gives detection limits in the tens of ppt range.

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References

Akyüz, M.: Simultaneous determination of aliphatic and aromatic amines in ambient air and airborne particulate matters by gas chromatography-mass spectrometry, *Atmos. Environ.*, 42, 3809–3819, 2008.

20 Almeida, J., Schobesberger, S., Kürten, A., Ortega, I. K., Kupiainen-Määttä, O., Praplan, A. P., Adamov, A., Amorim, A., Bianchi, F., Breitenlechner, M., David, A., Dommen, J., Donahue, N. M., Downard, A., Dunne, E., Duplissy, J., Ehrhart, S., Flagan, R. C., Franchin, A., Guida, R., Hakala, J., Hansel, A., Heinritzi, M., Henschel, H., Jokinen, T., Junninen, H., Kajos, M., Kangasluoma, J., Keskinen, H., Kupc, A., Kurtén, T., Kvashin, A. N., Laaksonen, A., Lehtipalo, K., Leiminger, M., Leppä, J., Loukonen, V., Makhmutov, V., Mathot, S., McGrath, M. J., Nieminen, T., Olenius, T., Onnela, A., Petäjä, T., Riccobono, F., Riipinen, I., Rissanen, M., Rondo, L., Ruuskanen, T., Santos, F. D., Sarnela, N., Schallhart, S., Schnitzhofer, R., Seinfeld, J. H., Simon, M., Sipilä, M., Stozhkov, Y., Stratmann, F., Tomé, A.,



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5 Anderson, N., Strader, R., and Davidson, C.: Airborne reduced nitrogen: ammonia emissions from agriculture and other sources, *Environ. Int.*, 29, 277–286, 2003.

10 Angelino, S., Suess, D., and Prather, K.: Formation of aerosol particles from reactions of secondary and tertiary alkylamines: characterization by aerosol time-of-flight mass spectrometry, *Environ. Sci. Technol.*, 35, 3130–3138, 2001.

Barsanti, K. C., McMurry, P. H., and Smith, J. N.: The potential contribution of organic salts to new particle growth, *Atmos. Chem. Phys.*, 9, 2949–2957, doi:10.5194/acp-9-2949-2009, 2009.

15 Bates, T., Lamb, B., Guenther, A., Dignon, J., and Stoiber, R.: Sulfur emissions to the atmosphere from natural sources, *J. Atmos. Chem.*, 14, 315–337, 1992.

Berndt, T., Stratmann, F., Sipilä, M., Vanhanen, J., Petäjä, T., Mikkilä, J., Grüner, A., Spindler, G., Lee Mauldin III, R., Curtius, J., Kulmala, M., and Heintzenberg, J.: Laboratory study on new particle formation from the reaction OH + SO₂: influence of experimental conditions, H₂O vapour, NH₃ and the amine tert-butylamine on the overall process, *Atmos. Chem. Phys.*, 10, 7101–7116, doi:10.5194/acp-10-7101-2010, 2010.

Borduas, N., Abbatt, J. P. D., and Murphy, J. G.: Gas phase oxidation of monoethanolamine (MEA) with OH radical and ozone: kinetics, products, and particles, *Environ. Sci. Technol.*, 47, 6377–6383, 2013.

25 Bzdek, B. R. and Johnston, M. V.: New particle formation and growth in the troposphere, *Anal. Chem.*, 82, 7871–7878, 2010.

Bzdek, B. R., Ridge, D. P., and Johnston, M. V.: Amine exchange into ammonium bisulfate and ammonium nitrate nuclei, *Atmos. Chem. Phys.*, 10, 3495–3503, doi:10.5194/acp-10-3495-2010, 2010.

30 Bzdek, B. R., Ridge, D. P., and Johnston, M. V.: Reactivity of methanesulfonic acid salt clusters relevant to marine air, *J. Geophys. Res.-Atmos.*, 116, D03301, doi:10.1029/2010JD015217, 2011.

Chan, L. P. and Chan, C. K.: Displacement of ammonium from aerosol particles by uptake of triethylamine, *Aerosol Sci. Tech.*, 46, 236–247, 2012.

Chen, M., Titcombe, M., Jiang, J., Jen, C., Kuang, C., Fischer, M. L., Eisele, F. L., Siepmann, J. I., Hanson, D. R., Zhao, J., and McMurry, P. H.: Acid–base chemical reaction model for nucleation rates in the polluted atmospheric boundary layer, *P. Natl. Acad. Sci. USA*, 109, 18713–18718, 2012.

5 Creamean, J. M., Ault, A. P., Ten Hoeve, J. E., Jacobson, M. Z., Roberts, G. C., and Prather, K. A.: Measurements of aerosol chemistry during new particle formation events at a remote rural mountain site, *Environ. Sci. Technol.*, 45, 8208–8216, 2011.

10 Dawson, M. L., Varner, M. E., Perraud, V., Ezell, M. J., Gerber, R. B., and Finlayson-Pitts, B. J.: Simplified mechanism for new particle formation from methanesulfonic acid, amines, and water via experiments and ab initio calculations, *P. Natl. Acad. Sci. USA*, 109, 18719–18724, 2012.

Feilberg, A., Liu, D., Adamsen, A. P. S., Hansen, M. J., and Jonassen, K. E. N.: Odorant emissions from intensive pig production measured by online proton-transfer-reaction mass spectrometry, *Environ. Sci. Technol.*, 44, 5894–5900, 2010.

15 Finlayson-Pitts, B. J. and Pitts, J. N.: Chemistry of the Upper and Lower Atmosphere – Theory, Experiments and Applications, Academic Press, San Diego, 2000.

Fournier, M., Lesage, J., Ostiguy, C., and Van Tra, H.: Sampling and analytical methodology development for the determination of primary and secondary low molecular weight amines in ambient air, *J. Environ. Monitor.*, 10, 379–386, 2008.

20 Fujii, T. and Kitai, T.: Determination of trace levels of trimethylamine in air by gas chromatography/surface ionization organic mass spectrometry, *Anal. Chem.*, 59, 379–382, 1987.

Fuselli, S., Benedetti, G., and Mastrangeli, R.: Determination of methylamines in air using activated-charcoal traps and gas-chromatographic analysis with an alkali flame detector (AFD), *Atmos. Environ.*, 16, 2943–2946, 1982.

25 Ge, X., Wexler, A., and Clegg, S.: Atmospheric amines – Part I. A review, *Atmos. Environ.*, 45, 524–546, 2011a.

Ge, X., Wexler, A., and Clegg, S.: Atmospheric amines – Part II. Thermodynamic properties and gas/particle partitioning, *Atmos. Environ.*, 45(3), 561–577, 2011b.

30 Gibb, S. W., Mantoura, R. F. C., and Liss, P. S.: Ocean–atmosphere exchange and atmospheric speciation of ammonia and methylamines in the region of the NW Arabian Sea, *Global Biogeochem. Cy.*, 13, 161–177, 1999.

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Gronberg, L., Lovkvist, P., and Jonsson, J.: Determination of aliphatic-amines in air by membrane enrichment directly coupled to a gas-chromatograph, *Chromatographia*, 33, 77–82, 1992.

Hansen, M. J., Adamsen, A. P. S., and Feilberg, A.: Recovery of odorants from an olfactometer measured by proton-transfer-reaction mass spectrometry, *Sensors*, 13, 7860–7871, 2013.

Hanson, D. R., McMurry, P. H., Jiang, J., Tanner, D., and Huey, L. G.: Ambient pressure proton transfer mass spectrometry: detection of amines and ammonia, *Environ. Sci. Technol.*, 45, 8881–8888, 2011.

Haynes, W. M. (Ed.): *Dissociation Constants of Organic Acids and Bases*, CRC Handbook of Chemistry and Physics, available at: <http://www.hbcpnetbase.com/> (last access: 29 April 2013), 2013.

Hiranuma, N., Brooks, S. D., Thornton, D. C. O., and Auvermann, B. W.: Atmospheric ammonia mixing ratios at an open-air cattle feeding facility, *J. Air Waste Manage.*, 60, 210–218, 2010.

Huang, G., Hou, J., and Zhou, X.: A measurement method for atmospheric ammonia and primary amines based on aqueous sampling, OPA derivatization and HPLC analysis, *Environ. Sci. Technol.*, 43, 5851–5856, 2009.

Kallinger, G. and Niessner, R.: Laboratory investigation of annular denuders as sampling system for the determination of aliphatic primary and secondary amines in stack gas, *Mikrochim. Acta*, 130, 309–316, 1999.

Kuhn, U., Sintermann, J., Spirig, C., Jocher, M., Ammann, C., and Neftel, A.: Basic biogenic aerosol precursors: agricultural source attribution of volatile amines revised, *Geophys. Res. Lett.*, 38, L16811, doi:10.1029/2011GL047958, 2011.

Kunin, R. and Barry, R. E.: Carboxylic, weak acid type, cation exchange resin, *Ind. Eng. Chem.*, 41, 1269–1272, 1949.

Kuwata, K., Akiyama, E., Yamazaki, Y., Yamasaki, H., and Kuge, Y.: Trace determination of low molecular weight aliphatic amines in air by gas chromatograph, *Anal. Chem.*, 55, 2199–2201, 1983.

Liu, D., Feilberg, A., Adamsen, A. P. S., and Jonassen, K. E. N.: The effect of slurry treatment including ozonation on odorant reduction measured by in-situ PTR-MS, *Atmos. Environ.*, 45, 3786–3793, 2011.

Liu, Y., Han, C., Liu, C., Ma, J., Ma, Q., and He, H.: Differences in the reactivity of ammonium salts with methylamine, *Atmos. Chem. Phys.*, 12, 4855–4865, doi:10.5194/acp-12-4855-2012, 2012.

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Lloyd, J., Heaton, K., and Johnston, M.: Reactive uptake of trimethylamine into ammonium nitrate particles, *J. Phys. Chem. A*, 113, 4840–4843, 2009.

Lobert, J. M., Scharffe, D. H., Hao, W. M., and Crutzen, P. J.: Importance of biomass burning in the atmospheric budgets of nitrogen-containing gases, *Nature*, 346, 552–554, 1990.

5 Loukonen, V., Kurtén, T., Ortega, I. K., Vehkamäki, H., Pádua, A. A. H., Sellegrí, K., and Kulmala, M.: Enhancing effect of dimethylamine in sulfuric acid nucleation in the presence of water – a computational study, *Atmos. Chem. Phys.*, 10, 4961–4974, doi:10.5194/acp-10-4961-2010, 2010.

10 Mosier, A., Andre, C., and Viets, F.: Identification of aliphatic amines volatilized from cattle feedyard, *Environ. Sci. Technol.*, 7, 642–644, 1973.

Müller, C., Iinuma, Y., Karstensen, J., van Pinxteren, D., Lehmann, S., Gnauk, T., and Herrmann, H.: Seasonal variation of aliphatic amines in marine sub-micrometer particles at the Cape Verde islands, *Atmos. Chem. Phys.*, 9, 9587–9597, doi:10.5194/acp-9-9587-2009, 2009.

15 Murphy, S. M., Sorooshian, A., Kroll, J. H., Ng, N. L., Chhabra, P., Tong, C., Surratt, J. D., Knipping, E., Flagan, R. C., and Seinfeld, J. H.: Secondary aerosol formation from atmospheric reactions of aliphatic amines, *Atmos. Chem. Phys.*, 7, 2313–2337, doi:10.5194/acp-7-2313-2007, 2007.

Ni, J.-Q., Robarge, W. P., Xiao, C., and Heber, A. J.: Volatile organic compounds at swine 20 facilities: a critical review, *Chemosphere*, 89, 769–788, 2012.

Nielsen, C. J., Herrmann, H., and Weller, C.: Atmospheric chemistry and environmental impact of the use of amines in carbon capture and storage (CCS), *Chem. Soc. Rev.*, 41, 6684–6704, 2012.

25 Nishikawa, T. and Kuwata, K.: Liquid chromatographic determination of low molecular weight aliphatic amines in air via derivatization with 7-chloro-4-nitro-2,1,3-benzoxadiazole, *Anal. Chem.*, 56, 1790–1793, 1984.

Nishino, N., Arquero, K. D., Dawson, M. L., and Finlayson-Pitts, B. J.: Infrared studies of the reaction of methanesulfonic acid with trimethylamine on surfaces, *Environ. Sci. Technol.*, 48, 323–330, 2013.

30 NOAA: Quality Controlled Local Climatological Data, NOAA, available at: <http://cdo.ncdc.noaa.gov/qclcd/QCLCD> (last access: 6 January 2014), 2014.

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Orsini, D. A., Ma, Y., Sullivan, A., Sierau, B., Baumann, K., and Weber, R. J.: Refinements to the particle-into-liquid sampler (PILS) for ground and airborne measurements of water soluble aerosol composition, *Atmos. Environ.*, 37, 1243–1259, 2003.

Praplan, A. P., Bianchi, F., Dommen, J., and Baltensperger, U.: Dimethylamine and ammonia measurements with ion chromatography during the CLOUD4 campaign, *Atmos. Meas. Tech.*, 5, 2161–2167, doi:10.5194/amt-5-2161-2012, 2012.

Rochelle, G. T.: Amine scrubbing for CO₂ capture, *Science*, 325, 1652–1654, 2009.

Ruiz-Jiménez, J., Hautala, S., Parshintsev, J., Laitinen, T., Hartonen, K., Petäjä, T., Kulmala, M., and Riekola, M.-L.: Aliphatic and aromatic amines in atmospheric aerosol particles: comparison of three ionization techniques in liquid chromatography-mass spectrometry and method development, *Talanta*, 97, 55–62, 2012.

Santagati, N. A., Bousquet, E., Spadaro, A., and Ronsisvalle, G.: Analysis of aliphatic amines in air samples by HPLC with electrochemical detection, *J. Pharmaceut. Biomed.*, 29, 1105–1111, 2002.

Schade, G. and Crutzen, P.: Emission of aliphatic amines from animal husbandry and their reactions: potential source of N₂O and HCN, *J. Atmos. Chem.*, 22, 319–346, 1995.

Schreiber, A., Zapp, P., and Kuckshinrichs, W.: Environmental assessment of German electricity generation from coal-fired power plants with amine-based carbon capture, *Int. J. Life Cycle Ass.*, 14, 547–559, 2009.

Seinfeld, J. H. and Pandis, S. N.: *Atmospheric Chemistry and Physics: from Air Pollution to Climate Change*, Wiley-Interscience, 2006.

Smith, J., Barsanti, K., Friedli, H., Ehn, M., Kulmala, M., Collins, D., Scheckman, J., Williams, B., and McMurry, P.: Observations of aminium salts in atmospheric nanoparticles and possible climatic implications, *P. Natl. Acad. Sci. USA*, 107, 6634–6639, 2010.

Tang, X., Price, D., Praske, E., Lee, S. A., Shattuck, M. A., Purvis-Roberts, K., Silva, P. J., Asa-Awuku, A., and Cocker, D. R.: NO₃ radical, OH radical and O₃-initiated secondary aerosol formation from aliphatic amines, *Atmos. Environ.*, 72, 105–112, 2013.

Tanimoto, H., Aoki, N., Inomata, S., Hirokawa, J., and Sadanaga, Y.: Development of a PTR-TOFMS instrument for real-time measurements of volatile organic compounds in air, *Int. J. Mass Spectrom.*, 263, 1–11, 2007.

Toda, K., Koga, T., Tanaka, T., Ohira, S.-I., Berg, J. M., and Dasgupta, P. K.: Miniature open channel scrubbers for gas collection, *Talanta*, 82, 1870–1875, 2010.

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Table 1. Retention times and calculated detection limits for ammonia and amines. Errors shown are ± 2 s.

Species	Retention Time (min)	Liquid Phase Detection Limit ^a ($M \times 10^{-7}$)	Gas Phase Detection Limit ^{a, b}	
			High-Concentration Cartridge (ppb in air)	Low-Concentration Cartridge (ppt in air)
Ammonia	7.5	2.3 ± 1.6	2.8 ± 1.9	12 ± 8
Methylamine	8.8	2.6 ± 1.7	3.1 ± 2.1	14 ± 9
Dimethylamine	11.8	3.5 ± 2.4	4.3 ± 3.0	19 ± 13
Trimethylamine	20.7	8.2 ± 5.6	10 ± 7	45 ± 31

^a Detection limits are calculated from the average of the signal corresponding to 3/5 peak-to-peak noise from 10 cartridge samples.

^b Gas-phase detection limits for the high-concentration cartridge samples are based on 1 LPM sampling for 20 min, extraction in 10 mL oxalic acid, and injection of 20 μ L of the solution. For the low-concentration cartridge, detection limits are based on 150 $\text{cm}^3 \text{ min}^{-1}$ sampling for 60 min and in-line extraction on the IC.

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Table 2. Results of field measurements taken in Chino, CA along with weather data from NOAA^a for the Chino Airport.

Date (2013)	Start Time	Duration (min)	[NH ₃] (ppm)	[TMA] (ppb)	Temperature (°C)	Relative Humidity (%)
28 Aug	4:22 a.m.	45	0.90	6.8	21.1	57
28 Aug	5:08 a.m.	45	1.5	6.7	20.6	57 ^b
4 Sep	3:55 a.m.	50	0.75	4.0	21.1	71
4 Sep	4:47 a.m.	50	0.75	3.3	20.6	79
12 Sep	3:53 a.m.	50	0.19	1.3	15.0	93 ^c
12 Sep	4:45 a.m.	50	0.49	4.5	14.4	90 ^c

^a Available at: <http://cds.ncdc.noaa.gov/qclcd/QCLCD>^b Haze^c Mist

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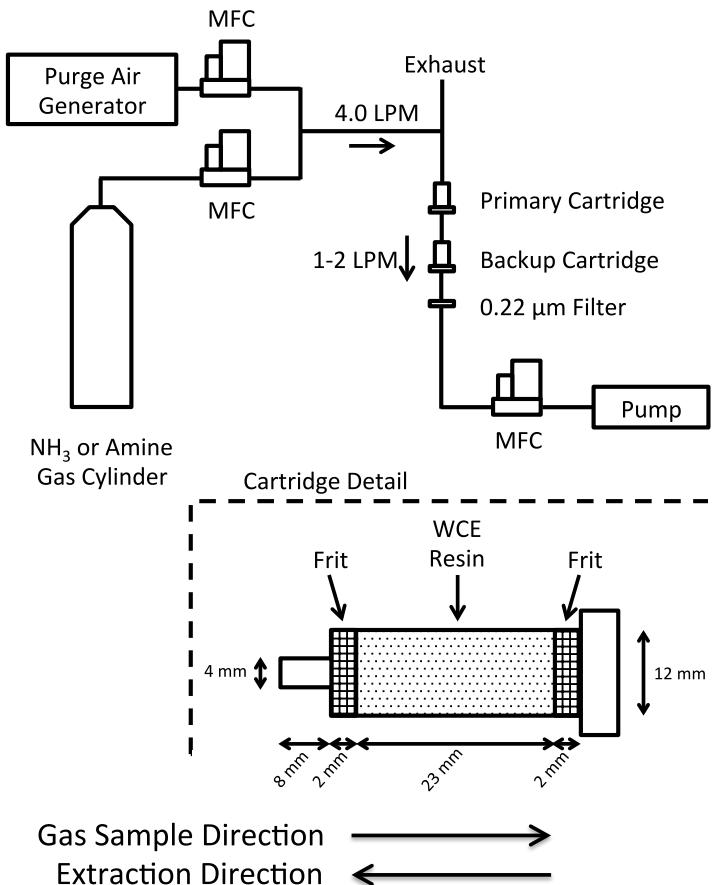


Fig. 1. Schematic of experimental system used to determine cartridge measurement efficiency. MFC = mass flow controller. Inset shows a detailed view of the “high-concentration” cartridge.

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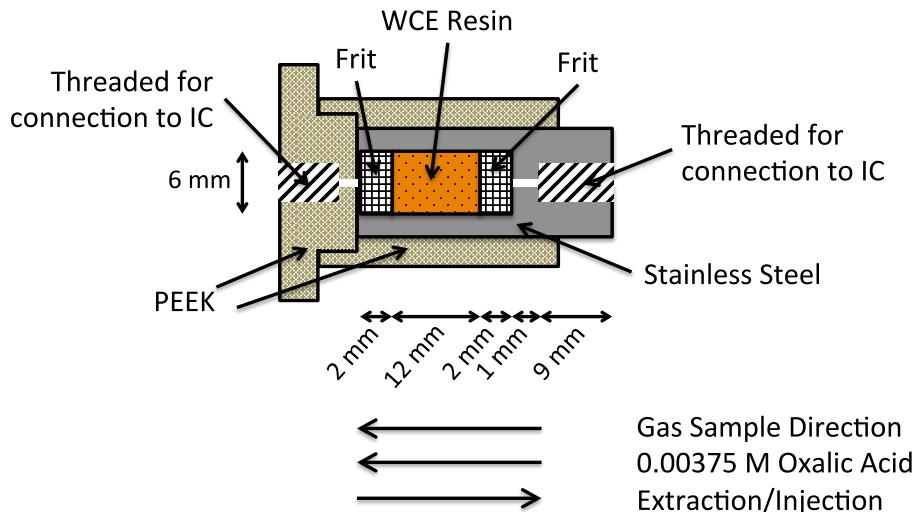


Fig. 2. Schematic of “low-concentration” cartridge.

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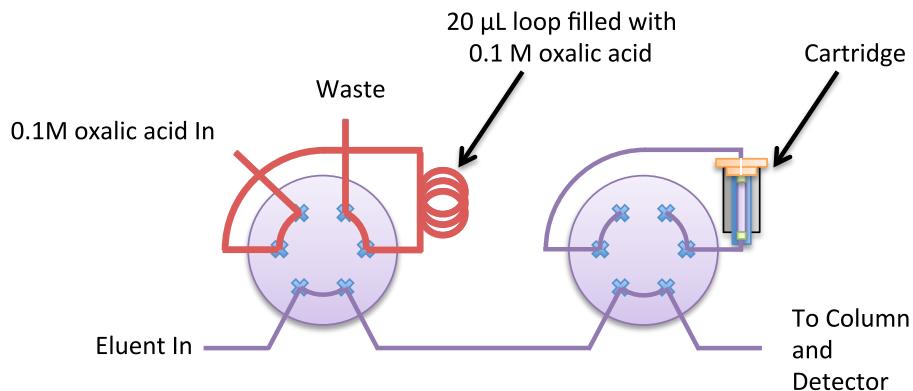


Fig. 3. Schematic of the in-line system for simultaneous extraction and analysis of ammonia and amine samples by IC (shown immediately prior to injection). At the beginning of the IC run, both injectors are actuated, allowing the 0.1 M oxalic acid plug to extract the cartridge and push the amine/ammonia onto the IC column.

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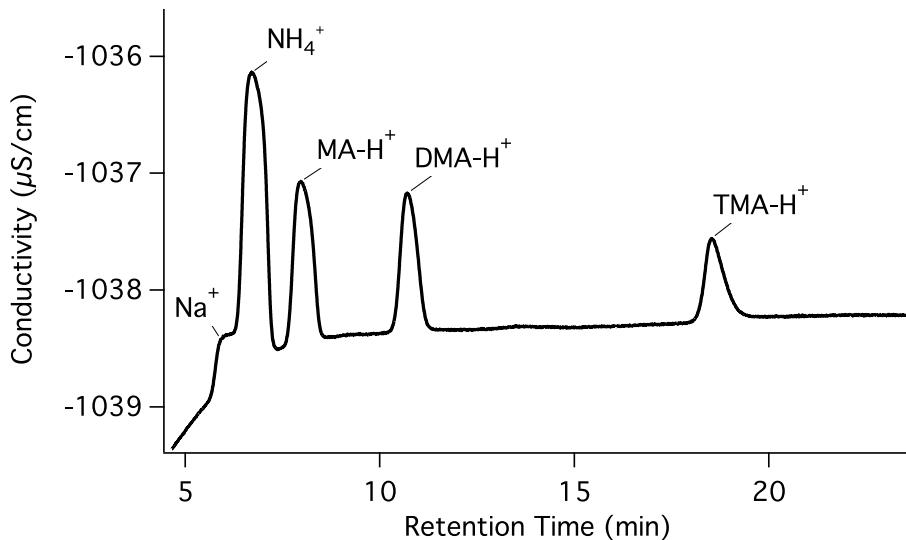


Fig. 4. A typical ion chromatogram for the amine/ammonia standards in 0.1 M oxalic acid. Standards included sodium methanesulfonate (NaCH_3SO_3 ; Aldrich; 98 %) because of the nature of ongoing laboratory experiments at the time so that Na^+ was also present.

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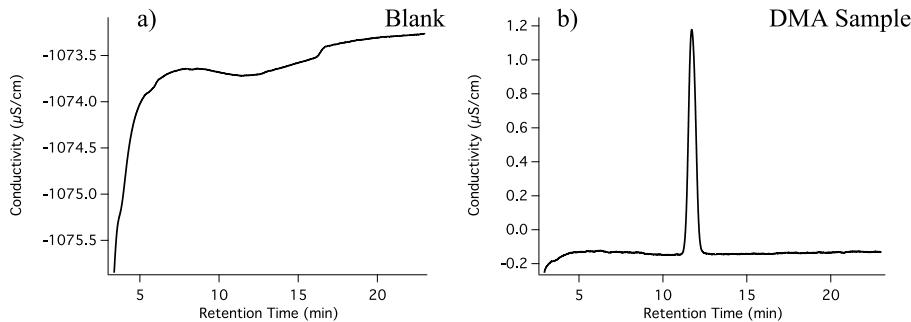


Fig. 5. (a) Chromatogram for a typical blank using high-concentration cartridges showing the characteristic baseline due to the high oxalic acid concentration in the cartridge extracts and **(b)** a background subtracted ion chromatogram for a DMA sample (nominally 1.0 ppm in N_2 ; see Sect. 3.1).

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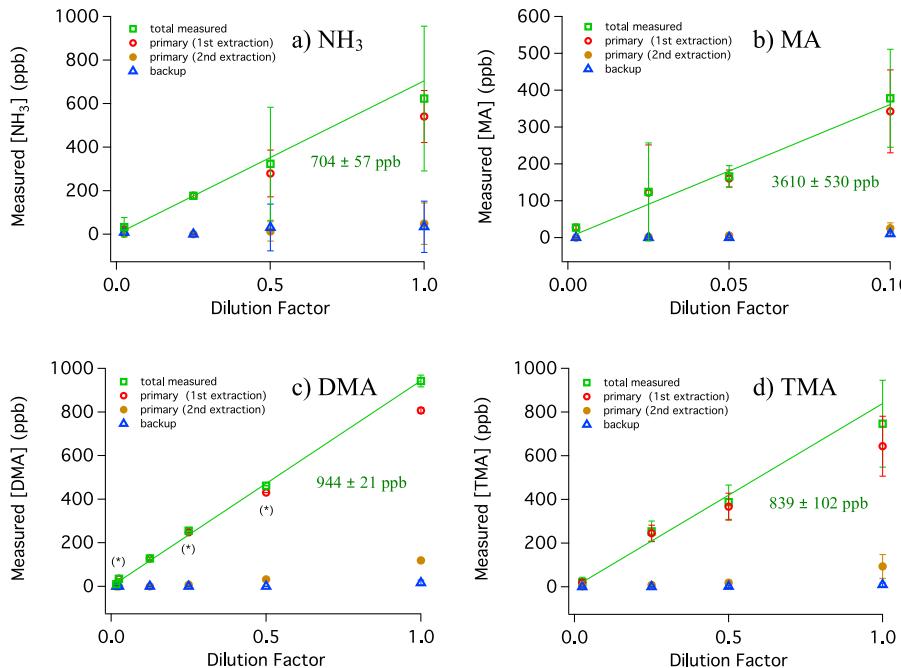


Fig. 6. Results for measurements of gas-phase standards of (a) ammonia, (b) MA, (c) DMA, and (d) TMA using high-concentration cartridges, including the first and second extract of the primary cartridge and the first extract of the backup cartridge, as well as the total measured concentrations. The dilution factor for ammonia or amine from the gas cylinders diluted with air is shown on the x axis, where 1.0 is the undiluted standard and 0.1 is a 10 % mixture. Data points marked with an asterisk (*) do not have sufficient replicates to include error bars. The green lines are weighted least-squares fits, where the weights for each point are given by $w = (1/s^2)$ and s is the sample standard deviation of the measurements at each dilution. Slopes of fitted lines are shown in green ($\pm 2s$). Manufacturer labeled concentrations for the undiluted standards were: 0.812 ppm NH₃, 10 ppm MA, 1.0 ppm DMA, and 1.0 ppm TMA.

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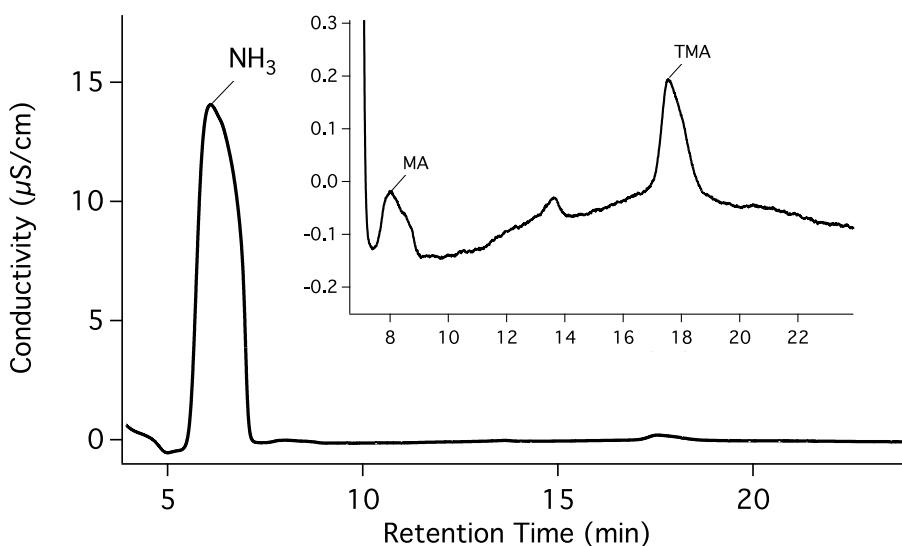


Fig. 7. The background-subtracted chromatogram from the first extract of the cartridge for the sample taken 28 August 2013 at 4:22 a.m. in Chino, CA. Inset: same chromatogram magnified to show peaks for MA and TMA.

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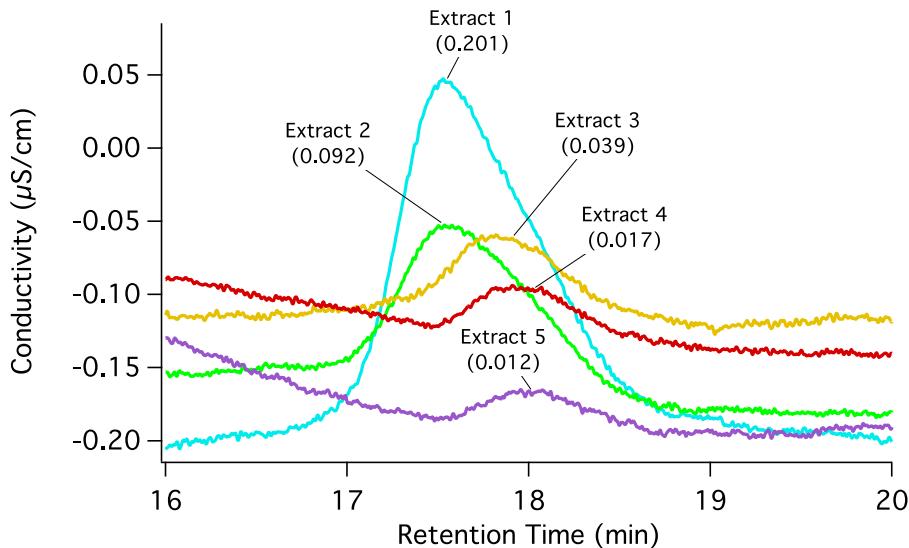


Fig. 8. The background-subtracted chromatograms from all five extracts of the cartridge for the sample taken 28 August 2013 at 5:08 a.m. in Chino, CA. Image is magnified to show TMA peak. Integrated peak areas in $(\mu\text{S}\text{min})\text{cm}^{-1}$ are shown in parentheses. The slight shift in retention time at lower peak size was typical for TMA in both standards and samples.