2) Page 5, line 2, Reaction R7:

Please change to:

“ 2NO2 + M 🡪 N2O4 + M, k7(298K, 105 Pa) ≈ 3.3×10-14 cm3 molec.-1 s-1 (R7)”

3) Page 5, line 4, Reaction R8: please change to:

“ N2O4 + M 🡪 2NO2 + M, k8(298K, 105 Pa) ≈ 1.47×105 s-1 (R8)”

6) Page 5, line 67, Reaction R9: please change to:

“ O+ O2 + M 🡪 O3 + M, k9(298, 105 Pa) ≈ 1.46×10-14 cm3 molec.-1 s-1 (R9) “

8) Page 7, line 4, Reaction R11: please change to:

“ O + NO2 + M 🡪 NO3 + M, k11(298, 105 Pa) ≈ 2.5 × 10-11 cm3 molec.-1 s-1 (R11) ”

9) Page 7, line 17, Reaction R15: please change to:

“NO3 + NO2 + M 🡪 N2O5 + M, k15(298K, 105 Pa) ≈ 1.34×10-12 cm3 molec.-1 s-1 (R15)”

Please change reaction equation R16 to:

“ N2O5 + M 🡪 NO3 + NO2 + M, k16(298K, 105 Pa) ≈ 0.07 s-1 (R16)”

11) Page 7, line 59, R19: Please change reaction equation R19 to:

“OH + NO2 + M 🡪 HNO3 + M, k19(298K, 105 Pa) ≈ 1.05 ⋅ 10-11 cm3 molec.-1 s-1 (R19)”

- Line labelled “(16)b,\*” (R16): The value in the rightmost column should read:

“0.07 s-1” instead of “2.98 s-1”

12) Page 9, Table 1:

- Line labelled “(16)b,\*” (R16): The value in the rightmost column should read:

“0.07 s-1” instead of “2.98 s-1”