Implementation of an IBBCEAS technique in an atmospheric simulation chamber for *in situ* NO$_3$ monitoring: characterization and validation for kinetic studies

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Abstract. An incoherent broadband cavity-enhanced absorption spectroscopy (IBBCEAS) technique has been developed for *in situ* monitoring of NO$_3$ radicals at the ppt level in the CSA simulation chamber (at LISA). The technique couples an incoherent broadband light source centered at 662 nm with a high finesse optical cavity made of two highly reflecting mirrors. The optical cavity which has an effective length of 82 cm allows for up to 3 km of effective absorption and a high sensitivity for NO$_3$ detection (up to 6 ppt for an integration time of 10 seconds). This technique also allows NO$_2$ monitoring (up to 9 ppb for an integration time of 10 seconds). Here, we present the experimental setup as well as tests for its characterization and validation. The validation tests include an intercomparison with another independent technique (FTIR) and the absolute rate determination for the reaction *trans*-2-butene + NO$_3$ which is already well documented in the literature. The value of $(4.13 \pm 0.45) \times 10^{-13}$ cm^3 molecule^{-1} s^{-1} has been found, which is in good agreement with previous determinations. From these experiments, optimal operation conditions are proposed. The technique is now fully operational and can be used to determine rate constants for fast reactions involving complex volatile organic compounds (with rate constants up to 10^{-10} cm^3 molecule^{-1} s^{-1}).

1. Introduction

The night time chemistry in polluted urban or sub-urban areas has been proved to be governed by NO$_3$ radicals since its discovery in the 1980s *(Naudet et al., 1981; Noxon et al., 1978, 1980; Platt et al., 1980)*. In particular, the NO$_3$ radical has been shown to be an efficient oxidant for some organic compounds, or in some cases even the dominant one, thus impacting the budget of these species and their degradation products. Unsaturated VOCs, including biogenic VOCs, are particularly reactive towards NO$_3$ radicals *(Wayne et al., 1991)*. Providing kinetic data for these reactions is essential for a better understanding of the role of NO$_3$ radicals in their degradation. Nevertheless, due to the high reactivity of some unsaturated VOCs with NO$_3$ (with rate constants which can reach 10^{-11} to 10^{-10} cm^3 molecule^{-1} s^{-1}), absolute rate determination for these reactions appears to be difficult as it requires the use of a highly sensitive method for NO$_3$ monitoring. As a consequence, the number of absolute kinetic studies for the NO$_3$-initiated oxidation of terpenes is very limited and this leads to large uncertainties on this chemistry as it has been pointed out in the literature *(Atkinson, 2000; Brown and Stutz, 2012; Ng et al., 2017)*. *Calvert et al., 2015* gave recommendations for NO$_3$ oxidation rate constants for 91 alkenes (ranging between 10^{-16} and 10^{-10} cm^3 molecule^{-1} s^{-1}) and more than 98 % of the determinations on which these recommendations are based were conducted using the relative rate method. One of the reasons for this is still the challenging measurement of NO$_3$

radicals at low mixing ratios (<100 ppt) during such experiments. For these compounds, new absolute determinations are essential to better evaluate the role of NO$_3$ radical in their degradation.

Among the various experimental tools which are currently used to measure rate constants, atmospheric simulation chambers represent suitable tools for performing experiments under very realistic atmospheric conditions. This implies low concentrations of reactants in order to minimize possible secondary reactions. Another benefit of these facilities is the high analytical capabilities which allow *in situ* monitoring of reactants and products with a high sensitivity. Even though significant progress has been made in the last decades for NO$_3$ radicals measurement at low concentrations with the arising of cavity enhanced and cavity ring down spectroscopic techniques (Ball et al., 2004; Bitter et al., 2005; Kennedy et al., 2011; Langridge et al., 2008) as well as laser induced fluorescence techniques (Matsumoto et al., 2005b, 2005c, 2005a; Wood et al., 2003), it is observed only few were coupled to simulation chambers (Dorn et al., 2013; Venables et al., 2006; Wu et al., 2014).In addition, to our knowledge, none of these techniques has been used for kinetic applications involving NO$_3$ radical in simulation chambers.

For this purpose, the analytical capabilities of the CSA chamber available at LISA have been improved by developing a sensitive technique for measuring NO$_3$ radicals at very low concentration. An incoherent broadband cavity enhanced absorption spectroscopy (IBBCEAS) technique has been coupled to the chamber with the objective of performing high sensitivity *in situ* NO$_3$ monitoring with integration time of seconds.

In this paper, we describe the experimental setup and the characterization of the technique. Finally, the IBBCEAS technique has been validated thanks to an intercomparison of NO$_2$ and NO$_3$ measurement with FTIR technique and an absolute rate determination for the well-known reaction *trans*-2-butene+NO$_3$.

2. Experimental section

2.1. The CSA chamber

The CSA chamber is made of a large and evacuable Pyrex® reactor (6 m length, 45 cm diameter and 977 L volume) which has been previously presented in detail (Doussin et al., 1997). It is equipped with a homogenization system which is made of i) an injection pipe (4 meters long, 1 cm diameter and regularly drilled with 1 mm holes) which allows to inject the reactants all along the chamber, ii) 2 stainless steel fans allowing a homogenization of gas inside the chamber and iii) a close-circuit Teflon pump connected at both ends allowing a recirculation of the gas mixing. This system allows a mixing time below one minute. The chamber is also equipped with two *in situ* spectroscopic analytical devices coupled with "White" type multiple reflection systems inside the reactor: (i) a FTIR spectrometer (Bruker Vertex 80) which allows acquiring spectra in the range of 600-4000 cm^{-1} with a maximal spectral resolution of 0.07 cm^{-1} and an optical path length of 204 m; (ii) an UV-Visible grating spectrometer consisting of a high-pressure Xenon arc lamp (Osram XBO, 450 W Xe UV), a monochromator HR 320 (Jobin-Yvon) and a CCD camera (CCD 3000, 1024 x 58 pixel, Jobin-Yvon) as a detector. This spectrometer allows acquiring spectra with a spectral resolution of 0.18 nm and an optical path length of 72 m.

This facility has intensively been used to investigate complex gas-phase chemistry involving organic compounds and to provide kinetic and mechanistic data. In particular, it has been used for absolute rate determination of reactions involving NO$_3$ radicals with a series of VOCs like ethers, esters and aldehydes (Kerdouci et al., 2012; Picquet-Varrault et al., 2009; Scarfogliero et al., 2006). In these studies, NO$_3$ was monitored at 662 nm with the *in situ* UV-visible spectrometer. However, due to the poor detection limit (0.5 ppb for 1 minute of acquisition),

and taking into account the experimental conditions, the range of rate constants that can be investigated is limited ($< 10^{-12}$ cm^3 molecule^{-1} s^{-1}) preventing from studying very reactive chemical systems, such as BVOC+NO$_3$ reactions.

2.2. The IBBCEAS setup

In order to improve the analytical capabilities of the CSA chamber, an IBBCEAS has been developed and coupled to the chamber for high sensitivity *in situ* NO$_3$ monitoring. A detailed description of the technique has been provided in previous works (Langridge et al., 2008; Romanini et al., 1997). IBBCEAS measurements are conducted by exciting with an incoherent broad-band source, a high finesse optical cavity formed by two mirrors with high reflectivity (R(λ) ~99.98 %). Photons resonate between the two mirrors increasing their lifetime in the cavity by a factor of 1/(1-R(λ)). During this time, photons traverse an effective path length of kilometers inside the cavity making possible observations of absorbing species at very low concentrations. The intensity transmitted by the optical cavity rapidly reaches steady state. The optical intracavity absorption coefficient of the sample $\alpha(\lambda)$ can then be calculated with the following expression if accurate measurement of cavity reflectivity R(λ) and of the distance between the mirrors (d) is provided (Engel et al., 1998):

$$\alpha(\lambda) = \left(\frac{I_0(\lambda)}{I(\lambda)} - 1\right)\left(\frac{1-R(\lambda)}{d}\right) \tag{Eq. 1}$$

Here I(λ) and I$_0$(λ) are the transmitted intensities measured in the presence and in absence of the absorbing species respectively.

The concentrations of the absorbing molecules can then be calculated using a least square algorithm to simultaneously fit the molecules absorption cross section using the formula:

$$\alpha(\lambda) = \sum_i [X_i]\, \sigma_{X_i}(\lambda) + p(\lambda) \tag{Eq. 2}$$

Where, $\sigma_{X_i}(\lambda)$ are the absorption cross sections, [X] are the species absorbing in the considered spectral region and p(λ) is a cubic polynomial to correct baseline deformations due to potential variations of the source intensity (Venables et al., 2006) or to absorption and/or scattering of particles in the chamber (Varma et al., 2013).

The optical cavity is made of two high reflectivity mirrors (Layertec, planar/concave mirrors with a 1 m radius of curvature, nominal reflectivity of 99.98 ± 0.01 % between 630 and 690 nm). It has been transversally installed on the CSA chamber using two co-axial outputs of the reactor. A scheme of the IBBCEAS instrument interfaced to the chamber is shown in Fig. 1. The distance between the mirrors is 82 cm and includes 45 cm for the chamber diameter, 2×10.5 cm for the Pyrex outputs and 2×8 cm for the interface mounts between the chamber and the commercial CRD Optics mount support. In order to prevent adsorption of semi-volatile species or deposition of particles on the mirrors, which would result in a significant decrease of the reflectivity, the mirrors can be protected thanks to a continuous nitrogen flush (flow rate: 300 mL min^{-1}) using a 1/16 inch input disposed close to the mirror surface. This flow rate has been optimized in order to efficiently protect the mirrors while limiting the dilution of the mixture in the measurement area. By comparing the absorption coefficients measured with and without the flush, for a known quantity of NO$_2$ in the chamber, this effective length was estimated to be d$_{eff}$ = 62 ± 3 cm (i.e. 24% lower than the physical length of the cavity). Nitrogen flush was not used in this study, but is available for

Thanks to the mixing system which ensures a fast homogenization of the mixture in the chamber, this effective length was observed to be constant during the whole duration of an experiment.
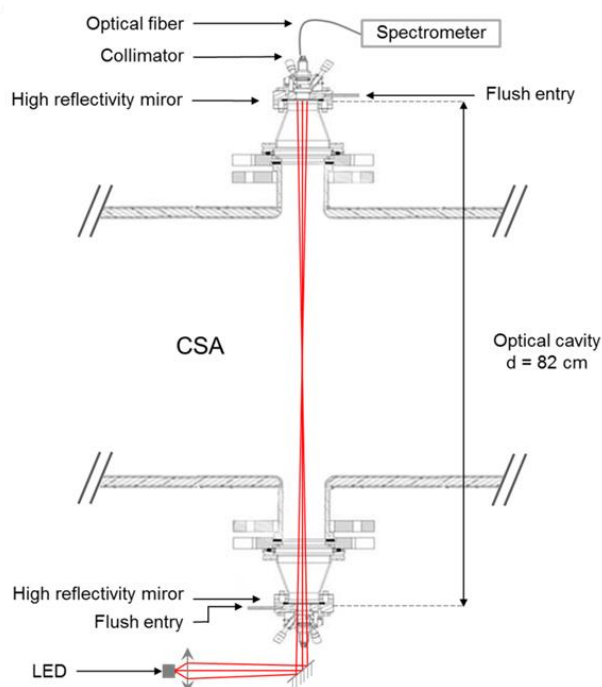
115 **Figure 1 : Scheme (transverse section) of the IBBCEAS instrument interfaced to the CSA chamber. Collimating lenses and curved mirrors are not shown on the beam injection side.**

A light emitting diode (LED, Mouser Electronics – Starboard, Luminus SST-10-DR-B130 DEEp Red, K, max. 430 mW, D5) with an approximative Gaussian shaped emission of 19 nm full width at half maximum (FWHM) and centered at 662 nm was used in order to monitor NO_3 at its maximum absorption wavelength. The LED

120 emission spectrum is compared to the cross section spectra of NO_3 (Orphal et al., 2003) and NO_2 (Vandaele et al., 1997) in Fig. 2. The spectral range of the LED is large enough to allow monitoring both NO_2 and NO_3.
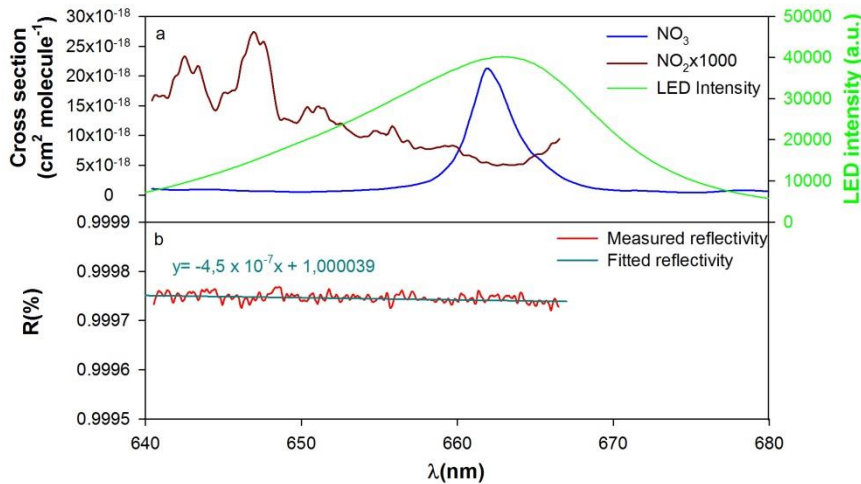
The LED is mounted on a thermo electric controller (TEC) device (ThermoElectric Cooling, Laser Mount Arroyo Instruments) to ensure a very precise temperature regulation ($\pm$ 0.01 °C) and stabilize the spectral distribution of the LED. With this device, changes in the LED intensity have been observed to be below 0.3 %. A laser diode controller (Arroyo Instrument 6310) provides the electric power for both TEC and LED (LED current is fixed at 900 mA). Light emitted by the LED is spatially incoherent and collimation is required for an effective coupling with the optical cavity. The light is hence focused with a convex lens (Thorlabs Aspheric Condenser Lens, 25.4 mm diameter, F = 16 mm, NA = 0.79) and injected into the optical cavity with two concave mirrors (Thorlabs, protected silver, diameter 50.8 mm, f = 50mm and diameter 75 mm, f = 500 mm respectively) in order to focus the beam in the middle of the cavity. Light transmitted through the cavity is directed thanks to a collimator (Thorlabs SMA Fiber Collimation Pkg, 635 nm, f = 35.41 mm, NA = 0.25) and an optical fiber (Ocean Optics VIS-NIR (200 μm slit, 5m long) to a miniature Ocean Optics QE65000 spectrometer. The spectrometer measures the cavity output wavelength dependent intensity and comprises a spectrograph interfaced to a charged coupled device (CCD) thermally stabilized at -15°C to minimize dark current. The spectral range covered by the spectrometer is 45 nm (640-685 nm) with a spectral resolution of 0.2 nm. In order to calculate the concentrations of absorbing species, a data treatment program has been developed in R (Ihaka and Gentleman, 1996) using a least square algorithm. A third-degree polynomial function is used in the fit to take into account baseline deformation due to small changes in the source intensity. The concentrations of the absorbing species and the polynomial function are fitted by minimizing the RMSE (root mean square error). In practice, the optimization was run following a bound optimization by quadratic approximation (BObyQA) method (Powell, 2009). The iterative process to minimize the RMSE (between absorption coefficients from Eq. 1 and Eq.2) stops when none of the parameters vary more than 0.2 % between two successive iterations. In the studied spectral range, absorbing species are H_2O, NO_2 and NO_3. Absorption by water vapor may be very high under atmospheric conditions. Nevertheless, due to dry conditions used during the experiments (RH < 1 %), H_2O absorption was negligible here. The absorption cross sections used are provided by the literature (Orphal et al., 2003; Vandaele et al., 1997) and have been convoluted with the apparatus function of the instrument. Because NO_2 cross sections provided by

Vandaele et al., 1997 are measured up to 666.5 nm, the treatment has systematically been conducted up to this value. The cross sections used for the data treatment are presented in Fig. 2.

3. Technique qualification and characterization

Several experiments have been carried out to assess the stability, the accuracy and the detection limit of the
155 technique. First experiments have been conducted to test the optical stability and the influence of pressure variations on the device. These tests have shown that the instrument is very stable (variations < 1 %). Two aspects have been shown to be particularly critical for measurement with IBBCEAS technique: I_0 measurement and the determination of mirror reflectivity.

3.1. Determination of the cavity reflectivity

160 Having a precise knowledge of the wavelength dependent mirror reflectivity, $R(\lambda)$, is one of the most critical point of the IBBCEAS technique (Venables et al., 2006). Two different methods have been proposed for accurate determination of $R(\lambda)$: i) measurement of a known concentration of an absorbing species (Ventrillard-Courtillot et al., 2010), ii) measurement of the ring-down time in the empty cavity using a pulsed laser CRDS technique (Ball et al., 2004). The first method has been employed here. The absorbing species which has been chosen for the
165 experiments is NO_2 as it absorbs in the whole spectral region of the LED emission and its absorption cross sections are known with high accuracy. NO_2 concentrations in the simulation chamber were obtained from *in situ* FTIR measurements using IBI_{NO2} (1530–1680 cm^{-1}) = $(5.6 \pm 0.2) \times 10^{-17}$ cm molecule^{-1} (base e,). To retrieve the mirror reflectivity $R(\lambda)$ from NO_2 concentration, the following equation is used:

$$R(\lambda) = 1 - d \times \sigma_{NO_2}(\lambda) \times [NO_2] \times \left(\frac{I_0(\lambda)}{I(\lambda)} - 1\right)^{-1} \qquad \text{(Eq. 3)}$$

170 where $\sigma_{NO_2}(\lambda)$ is the NO_2 absorption cross section (Vandaele et al., 1997) and $[NO_2]$ is the concentration of NO_2 determined by FTIR. In order to reduce the uncertainty on the reflectivity determination and to compensate the weak cross sections of NO_2 in the 660–670 nm region, high concentrations (up to 800 ppb) were used. A plot showing the variation of the reflectivity in function of the wavelength is presented as an example in Fig. 2. Due to the NO_2 reference spectrum, the reflectivity is measured up to 666.5 nm. During this experiment, the reflectivity
175 was observed to vary between 99.975 at 640 nm and 99.974 % at 667 nm and this is in agreement with the reflectivity provided by the supplier (99.98 $\pm$ 0.01 %). It was found to have a slight dependence with wavelengths ($y = -4.5 \times 10^{-7}x + 1.000039$), which justifies that it is necessary to measure it on a wide wavelengths range. At 662 nm, which corresponds to the maximum of NO_3 absorption, the reflectivity was found to be 99.974 %. At this wavelength, the effective absorption path length estimated from Eq. 4 is found to be 3.15 km:

180 $$X(\lambda) = d/(1 - R(\lambda)) \qquad \text{(Eq. 4)}$$

In addition, it has been observed that the reflectivity of the cavity can significantly vary with the mirror cleanliness. As an example, successive experiments showed that reflectivity can vary at 662 nm from 99.974 % to 99.971 % from one experiment to another, leading to variations of the effective absorption path length of almost 12 %. Therefore, it is crucial to precisely determine the reflectivity prior to each experiment.

185 ### 3.2. $I_0(\lambda)$ measurement

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Previous studies (Fuchs et al., 2010; Kennedy et al., 2011; Ventrillard-Courtillot et al., 2010) have pointed out that the $I_0(\lambda)$ has to be periodically recorded during an experiment to ensure accurate measurement with the IBBCEAS technique. Indeed, changes in the lamp emission spectrum or poor optical stability may induce changes in the absorption coefficient and therefore generate errors in the quantification of the species. This fact may be an issue for experiments in simulation chambers as the $I_0(\lambda)$ can only be recorded before injecting the reactants and experiments can then last for several hours.

In order to evaluate the stability of the signal during a typical experiment and the uncertainty generated by the use of a unique $I_0(\lambda)$ on the quantification of the species, two types of experiments have been performed: first, the stability of the optical system has been verified. For this purpose, long term measurements of the signal have been performed, leading to variations lower than 0.3% and to very small baseline distortions. Then, to verify the impact of these variations on the quantification of absorbing species, experiments were conducted by injecting NO_2, which was observed to be stable in the chamber, and by monitoring its concentration with the IBBCEAS technique for several hours: after the chamber has been filled with a mixture of N_2/O_2 (80/20) at atmospheric pressure, the $I_0(\lambda)$ has been measured. Then, a concentration of NO_2 was introduced into the chamber (mixing ratios ranging between 100 ppb and 1 ppm, depending of the experiment) and the signal $I(\lambda)$ was measured. From these measurements, the absorption coefficient and the NO_2 concentration have been calculated and plotted as a function of time in Fig. 3, for a concentration of 1 ppm of NO_2. An increase up to 3 % in NO_2 concentration has been observed 2 hours after recording the $I_0(\lambda)$ due to the deviation of the baseline which is no longer well corrected by the polynomial function. These results suggest that the accuracy of the measurement is significantly reduced after 2 hours. The length of the experiments should therefore not exceed this duration. Above this limit the uncertainty due to a unique measurement of the $I_0(\lambda)$ can be considered as negligible.
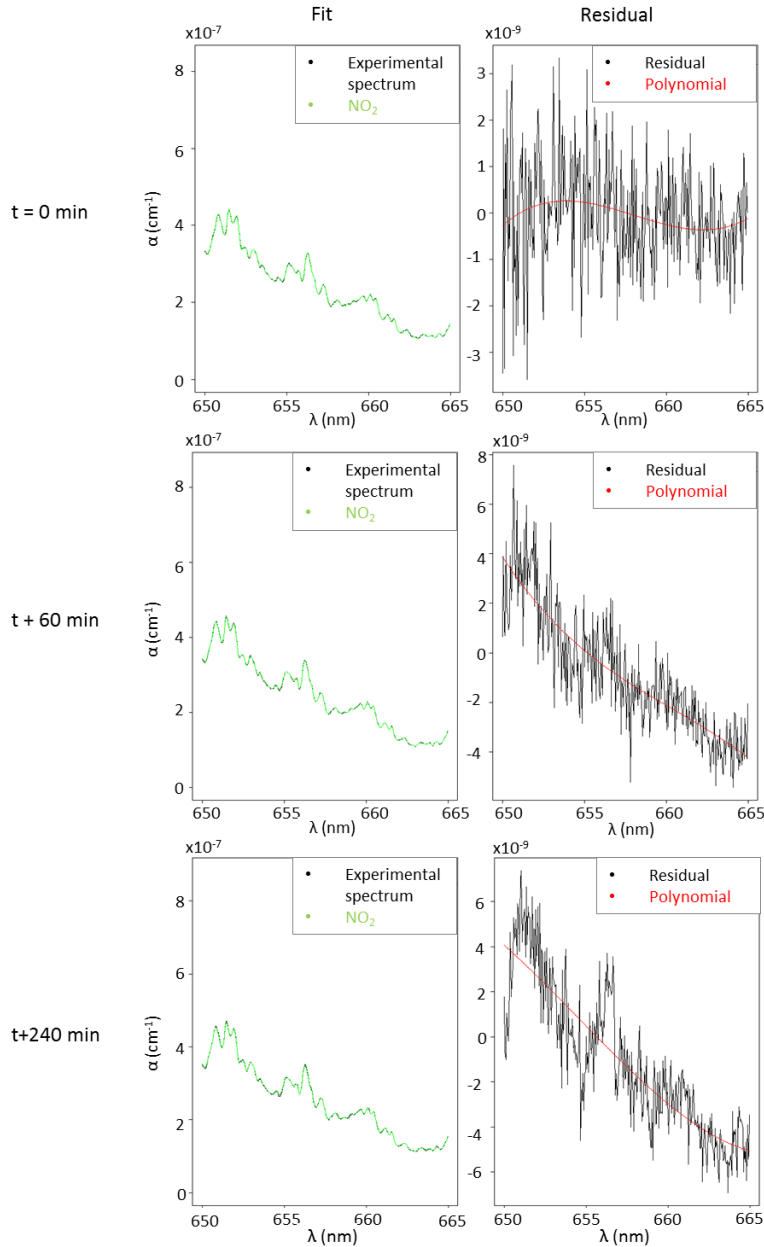
Figure 3 : Evolution of fit quality as a function of time since I_0 was made, measured for 1 ppm of NO_2

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3.3. Detection limit and Allan variance

The detection limit for NO_3 radical was estimated by considering 3 times the peak-to-peak noise on the absorption coefficient at 662 nm, which corresponds to the maximum absorption of NO_3 radical. For 10 seconds of acquisition time (which corresponds to 25 acquisitions of 400 ms) and an unpurged system, it has been found to be 1.2×10^{-9}

215   cm^{-1}. Considering that NO_3 radical cross section at this wavelength is 2.2×10^{-17} cm^2 $molecule^{-1}$, the detection limit for NO_3 was estimated to 6 ppt. The same approach was used to estimate the detection limit for NO_2. Between 645 and 650 nm, which correspond to the two main absorption peaks of NO_2, noise has been found to be 2×10^{-9} cm^{-1}. Considering the difference of maximum and minimum absorption of 2.2×10^{-20} cm^2 $molecule^{-1}$ in this range, detection limit has been found to be 11 ppb for 10s of integration time. A spectrum measured with an acquisition

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220 time of 10 s, for 6 ppt of NO_3 and 630 ppb of NO_2 is shown for illustration in Fig. 4. The used fit range was between 655 and 666.5 nm. Even at this low level of NO_3 concentration, the absorption is clear and allows its quantification. The fit appears to be satisfying for both NO_2 and NO_3 and the residual spectrum appears to be essentially made of noise, showing a good efficiency of the polynomial fit and a satisfying subtraction of species contributions. This figure shows that this wavelength range is efficient for a precise detection and quantification

225 of both species and validates the detection limit.
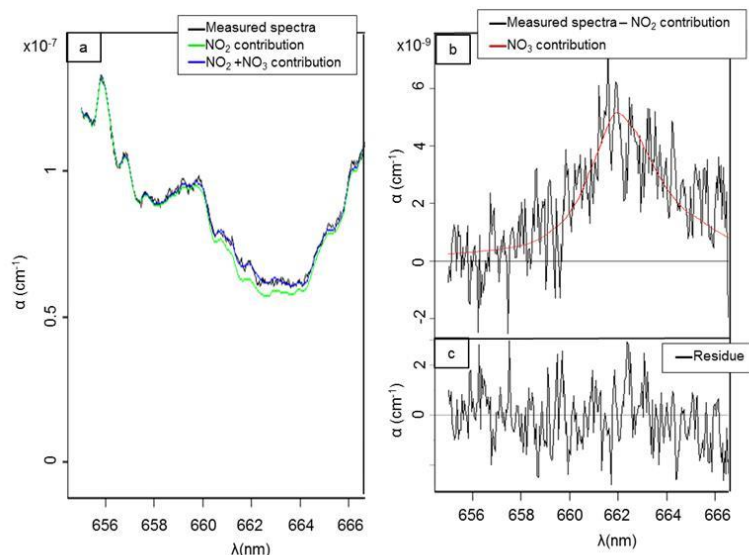
Figure 4: (a) measured absorption coefficient $\alpha(\lambda)$ (between 655 and 669 nm) with integration time of 10 s (in black); complete fit of NO_2 and NO_3 (in blue) with $[NO_3]$ = 6 ppt; $[NO_2]$ = 630 ppb; NO_2 fit only (in green) ; (b) measured absorption coefficient (in black) without NO_2 contribution and fitted with $[NO_3]$ = 6 ppt (in red); (c) residue of measured
230 **and fitted absorption coefficient.**

The potential of the IBBCEAS technique for measuring NO_3 radical during simulation chamber experiments has already been explored in previous works. It has been coupled to the simulation chamber at UCC (Cork, Ireland), to SAPHIR chamber at FZJ, (Jülich, Germany) and to CHARME chamber at LPCA (Wimereux, France). The characteristics and performances of the various instruments are compared in Table 1. Our instrument exhibits very

235 good performances with detection limit similar to those of the other instruments, but for shorter integration time and/or for smaller effective length. This reflects the very good stability of the optical system. These results also prove the potential of this technique for measuring NO_3 radical at low level of concentrations with a good time resolution (10 s) and thus suitable for kinetic measurements.

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Table 1 : Comparison of characteristics and performances of various IBBCEAS coupled to simulation chambers for the detection of NO₃ radicals.

In situ IBBCEAS	Deff* (cm)	DL / Integration time	Reference
LISA, Créteil France	**82**	**6 ppt / 10 s**	**Current work**
UCC, Cork, Ireland	462	4 ppt / 57 s	Venables et al., 2006
UCC, Jülich, Germany**	1800 ± 20	0.5-2 ppt/5 s	(Dorn et al., 2013)
LPCA, Wimereux, France	2000	7.9 ppt / 1 min	Wu et al., 2014

*Deff is the effective length of the cavity, calculated by taking into account the dilution generated by the mirror protective flush

**UCC's IBBCEAS was used on SAPHIR chamber during an intercomparison campaign.

In order to evaluate if our detection limit can be improved by increasing the integration time, the Allan variance has been calculated for various integration times during an experiment in which NO_2 concentration was monitored. The NO_2 mixing ratio was approximatively 1300 ppb. The Allan variance σ_A^2 is given by the equation:

$$\sigma_A^2 = \frac{1}{2(M-1)} \sum_{i=1}^{M-1} [x_{i+1}(t_{av}) - x_i(t_{av})]^2 \tag{Eq.5}$$

with M the number of measurements, t_{av} the integration time and x the concentration of NO_2 measured. In this experiment, 30 000 measurements of 2 seconds have been performed and the Allan variance was then calculated for various integration times ranging between 2 and 4096 seconds. The standard deviation of Allan, defined as the square root of the Allan variation provides an indication of the instrument stability in time. It is plotted as a function of the integration time in Fig. 5. For very short integration time (few seconds) the Allan deviation is very high due to the white noise of the instrument. The Allan deviation decreases with increasing integration time until 100 seconds. For longer integration times, the deviation increases with increasing integration time. Nevertheless, the deviation is low, showing that the instrument is very stable. Due to this stability, we expect that the detection limits calculated before can be improved by increasing integration time. From this test, it can be also concluded that the optimal integration time is around 100 seconds.
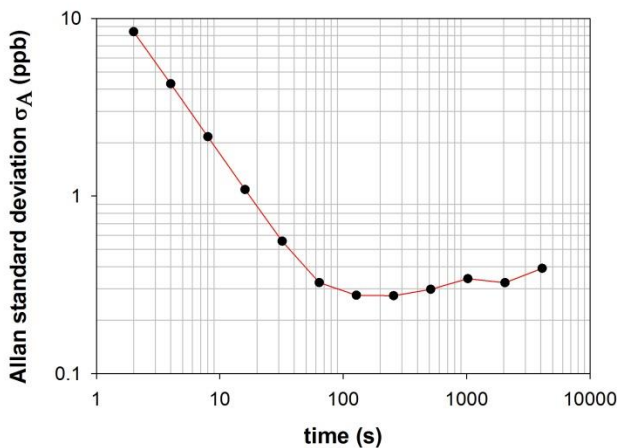

Figure 5: Allan variance vs. integration time calculated for the IBBCEAS technique

The high stability of the Allan deviation after 100 s also suggests that the stability of the optical device is optimal to perform measurements without recording a new $I_0(\lambda)$ for at least 4096 s, in agreement with the result of the test

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presented in section 3.2. In conclusion, the good stability of the optical device complies with the constraint of experiments in simulation chamber where $I_0(\lambda)$ can only be recorded at the beginning of the experiment.

3.4. Determination of the uncertainty

270 Considering Eq.1 and Eq. 2, the overall relative error on NO_3 concentration can be considered as the square root of the sum of the square relative errors, the reflectivity $R(\lambda)$ and the NO_3 absorption cross sections. In case of the use of the nitrogen flush, the uncertainty on d_{eff} has to be also taken into account (8 %, see Sect. 2). Considering Eq. 3, the uncertainty on $R(\lambda)$ should include the uncertainty on NO_2 concentration measured by FTIR estimated to be 8 % as well as the uncertainty on NO_2 absorption cross sections. The uncertainties on NO_2 and NO_3 cross

275 sections are estimated to be 3 % in the spectral range of interest (Vandaele et al. 1997 and Orphal, Fellows and Flaud, 2003). However, the uncertainty generated by the data treatment Δfit, i.e. by the fit, which results mainly from the noise in the spectra, should also be taken into account. Because the nitrate radical is an unstable species, this uncertainty cannot be estimated by calculating the standard deviation on its concentrations measured for a long period of monitoring. It was therefore estimated by considering the noise of a NO$_3$ concentration time profile

280 and has been found to be ~ 3ppt for 10 seconds of integration time. The overall absolute error on NO_3 concentration is then expressed by the following formula:

$$\Delta N_{NO_3} = \sqrt{\left(\left(\frac{\Delta\sigma_{NO_2}(\lambda)}{\sigma_{NO_2}(\lambda)}\right)^2 + \left(\frac{\Delta N_{NO_2,refl}}{N_{NO_2,refl}}\right)^2 + \left(\frac{\Delta\sigma_{NO_3}}{\sigma_{NO_3}}\right)^2\right)} \times N_{NO_3} + \Delta fit \qquad \text{(Eq. 6)}$$

where $\frac{\Delta\sigma_{NO_2}(\lambda)}{\sigma_{NO_2}(\lambda)}, \frac{\Delta N_{NO_2,refl}}{N_{NO_2,refl}}, \frac{\Delta\sigma_{NO_3}}{\sigma_{NO_3}}$ are relative uncertainties on NO_2 cross sections, NO_2 concentration used for the reflectivity measurement, NO_3 cross sections respectively and N_{NO_3} the concentration of NO_3. For 10 seconds of

285 integration time, the uncertainty is thus 9 % with an absolute part of 3 ppt.

4. Intercomparison study

After having defined the optimal operation conditions of the IBBCEAS, the technique has been validated thanks to an intercomparison with another instrument. During a dedicated experiment, NO_3 and NO_2 concentrations were measured by the IBBCEAS technique while NO_2 and N_2O_5 were monitored by *in situ* FTIR. The chamber was

290 first filled with dry synthetic air (RH < 1%) at atmospheric pressure and $I_0(\lambda)$ spectra were recorded for both FTIR and IBBCEAS. Several hundred of ppb of NO_2 (Air Liquide N20, purity>99 %, H_2O<3000 ppm) were then introduced into the chamber and IBBCEAS spectra were recorded in order to determine the mirrors reflectivity $R(\lambda)$ (see Sect. 3.2). NO_3 radicals were then formed *in situ*, using thermal dissociation of N_2O_5 (R. (1)), which was synthesized in a vacuum line following the reaction between O_3 and NO_2 (R. (2) and R. (3)). This protocol was

295 adapted from (Atkinson et al., 1984; Schott and Davidson, 1958) and is detailed in Picquet-Varrault et al., 2009.

$$N_2O_5 + M \leftrightarrows NO_3 + NO_2 + M \qquad \text{(R. 1)}$$
$$O_3 + NO_2 \rightarrow NO_3 + O_2 \qquad \text{(R. 2)}$$
$$NO_3 + NO_2 + M \rightarrow N_2O_5 + M \qquad \text{(R. 3)}$$

Then, by considering the following equilibrium:

300 $$NO_2 + NO_3 + M \rightarrow N_2O_5 + M \ (k_1) \qquad \text{(R. 4)}$$

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$$N_2O_5 + M \rightarrow NO_2 + NO_3 + M \ (k_2) \tag{R. 5}$$

And by assuming that this equilibrium is reached, NO_3 concentration can be deduced from NO_2 and N_2O_5 concentrations measured by FTIR. This hypothesis appears justified as no other reactive species has been introduced into the chamber and may thus disturb the equilibrium. The equilibrium constant ($K = k_1/k_2$) is well known and has been shown to highly depend on temperature and pressure (Atkinson et al., 2004). These two parameters were therefore precisely monitored during the experiment allowing calculating an equilibrium constant of 2.17×10^{-11} cm^3.molecule^{-1} at 298K and at 1030 mbar, using IUPAC database parameters (Atkinson et al., 2004). IBBCEAS spectra were recorded every 30 seconds while FTIR ones were recorded every 2 minutes. The integrated band intensities used to quantify NO_2 and N_2O_5 with FTIR were: IBI_{NO2} (1530-1680 cm^{-1}) = (5.6 ± 0.2) $\times$ 10^{-17} cm molecule^{-1} and IBI_{N2O5} (1200-1285 cm^{-1}) = (4.05 ± 0.4) $\times$ 10^{-17} cm molecule^{-1} (base e). The correlation plots between FTIR and IBBCEAS for NO_3 and NO_2 measurements are shown in Fig. 6. NO_2 concentrations measured by the two techniques are in very good agreement (the maximum difference between the two techniques is 6 %), with a slope of 1.0 ± 0.2. Here, the overall uncertainty was calculated as the sum of the statistical error on the slope (twice the standard deviation, 4 %) and systematic errors on FTIR (i.e. on IBI_{NO2}, 4 %) and IBBCEAS measurements (which includes uncertainties on NO_2 cross sections and on the mirrors reflectivity, 9%). The intercept of the linear regression (b = 15.0 ppb) is not significantly different from zero as it is lower than its uncertainty calculated as twice the standard deviation ($\Delta b = 25.1$ ppb). For NO_3 radical, the concentrations obtained by the two techniques are also in good agreement for the whole range of concentrations, from few ppt to several hundred ppt. The slope of the NO_3 correlation plot is 1.1 ± 0.3 suggesting a bias of 10 % between the two techniques. The error is calculated with the same method as for NO_2. However, this difference is not significant in regards to the uncertainties which are represented by the black dashed lines. For FTIR measurement, the uncertainties are calculated being the error on NO_2 and N_2O_5 measurement and on the equilibrium constant (21 %). The calculation for uncertainty on NO_3 IBBCEAS measurement is presented in section 3.4. The intercept appears to be very low (around 3 ppt).
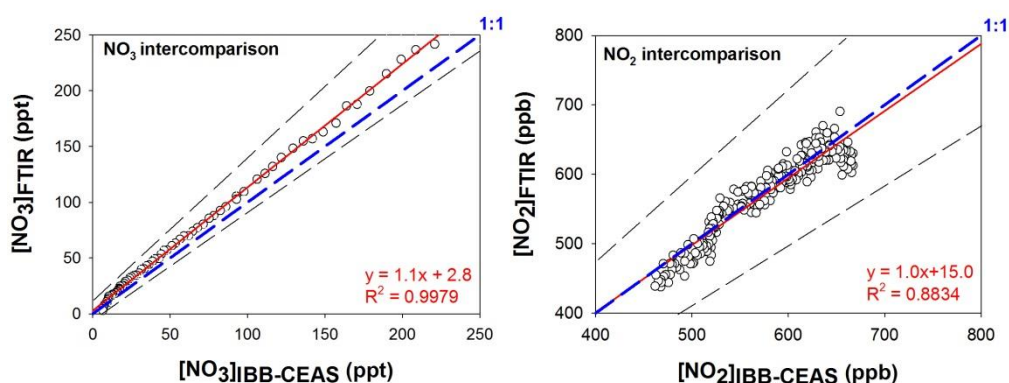

Figure 6 : Correlation between FTIR and IBBCEAS measurements for NO₃ (left graph) and NO₂ (right graph). Uncertainties are shown by dashed straight lines. Blue dashed lines show the 1:1 ratio.

In conclusion, the IBBCEAS exhibits very good agreement with the FTIR, for both NO_3 and NO_2 monitoring, with good sensitivity. This agreement is very satisfactory considering the fact that the IBBCEAS samples across the reactor width while the FTIR provides an integrated measurement on the whole reactor length. Finally, the intercomparison experiment shows that an eventual loss due to photolysis of NO_3 by the beam is not significant

because it would have led to lower concentrations measured by the IBBCEAS than by the FTIR. In addition, for wavelength longer than 640 nm, which is the case here, Johnston et al., 1996 have shown that photolysis quantum yield is close to 0. It is thus expected that the photolysis of NO_3 in the used wavelength range is not occuring.

335 Furthermore even though NO_3 was slightly subject to photolysis locally, the volume enlightened by the probe beam is very small in comparison to the overall volume of the chamber ($< 0.04\%$) and the homogenization system allowed a mixing with the rest of the volume.

5. Kinetic study: NO_3 + *trans*-2-butene

The last step of the validation consisted in a kinetic experiment in order to assess the potential of the technique for
340 kinetic studies: the IBBCEAS has been used to measure the rate constant of a well-known reaction: *trans*-2-butene + NO_3. This reaction has been chosen as it has been intensively studied in the literature. Six absolute rate determinations (Benter et al., 1992; Berndt et al., 1998; Dlugokencky and Howard, 1989; Kasyutich et al., 2002; Ravishankara and Mauldin, 1985; Rudich et al., 1996) and two relative ones (Atkinson et al., 1984; Japar and Niki, 1975) have been published leading to a recommendation by IUPAC (Atkinson et al., 2006). This reaction is also
345 particularly interesting because it does not produce SOA. This will allow us to test the performances of the instrument for monitoring NO_3 concentrations with a high time resolution and to validate our kinetic determination by comparison with previous ones.

The rate constant was determined using the absolute rate technique and by measuring the consumption of *trans*-2-butene due to its reaction with NO_3. Because no other oxidant was present in the mixture, it was therefore assumed
350 that *trans*-2-butene is consumed only by reaction with nitrate radical:

$$trans\text{-}2\text{-}butene + NO_3 \rightarrow Products \ (k_{trans\text{-}2\text{-}butene}) \tag{R. 6}$$

For this reaction, kinetic equation can be established as:

$$\frac{-d[trans\text{-}2\text{-}butene]}{dt} = k_{trans-2-butene}[NO_3][trans\text{-}2\text{-}butene] \tag{Eq. 7}$$

By making the hypothesis of small variations of time and [*trans*-2-butene], this relationship can be approximated
355 to:

$$-\Delta[trans\text{-}2\text{-}butene] = k_{trans-2-butene}[NO_3][trans\text{-}2\text{-}butene]\Delta t \tag{Eq. 8}$$

Where $\Delta[trans\text{-}2\text{-}butene]$ corresponds to the consumption of *trans*-2-butene during the time interval Δt and [*trans*-2-butene] and [NO_3] are averaged concentrations measured during this period. By plotting $-\Delta$[*trans*-2-butene] vs. the product [*trans*-2-butene][NO_3]Δt, a straight line with the slope corresponding to $k_{trans\text{-}2\text{-}butene}$ is
360 obtained. It is important to notice that because this absolute kinetic method consists of measuring the decay of *trans*-2-butene for a known concentration of NO_3, and not the decay of NO_3 radicals for a known concentration of the VOC, it is not affected by NO_3 additional loss processes (e.g. wall losses, reactions with NO_2 or with peroxy radicals). Only additional losses of the VOC would lead to an overestimation of the rate constant. So, it was checked prior to the experiments (i.e. before in the injection of N_2O_5) that no significant loss of the VOC was
365 observed in the timescale of the experiment (see below).

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Six kinetic experiments have been conducted in the dark, at room temperature (292 - 294 K) and atmospheric pressure in synthetic air. The initial conditions of reactants (*trans*-2-butene, N_2O_5, NO_2) are listed in Table 2. The reflectivity was measured prior to each experiment by introducing NO_2 into the chamber. When present, NO_2 initial concentrations were used also to slow down the reaction by shifting the N_2O_5 equilibrium. *Trans*-2-butene

370 (Air Liquide, purity>99 %) was then introduced and it was checked that no significant loss was observed in the timescale of the experiment. Nitrate radicals were generated into the simulation chamber from the thermal decomposition of dinitrogen pentoxide. N_2O_5 was stepwise injected in order to assure a consumption of *trans*-2-butene in a proper timescale to satisfactory monitor the reactants. Time resolved concentrations of *trans*-2-butene, NO_2 and N_2O_5 were monitored from their infrared absorption spectra every 2 minutes. The integrated band

375 intensity used to quantify the VOC is $IBI_{trans\text{-}2\text{-}butene}$ (870-1100 cm^{-1}) = (2.8 ± 0.3) × 10^{-18} cm molecule^{-1} (base e, measured by previous internal work). NO_3 was monitored with the IBBCEAS technique with an acquisition time of 30 seconds.

Table 2 : Injected mixing ratios for the kinetic study of the reaction *trans*-2-butene + NO_3

Experiment	$[NO_2]_0$ (ppb)	$[N_2O_5]$ (ppb) × number of injections	$[trans\text{-}2\text{-}butene]_0$ (ppb)
1	/	2500	1920
2	/	300 × 2 ; 150 × 2	750
3	920	1000 × 3	990
4	950	1500 × 2	1110
5	750	300 × 3	1110
6	/	300 × 3	1030

380 Figure 7. shows time profiles of reactants during a typical experiment. At the moment when N_2O_5 is injected, a rapid decrease of *trans*-2-butene and NO_3 concentrations is observed together with a large production of NO_2 due to N_2O_5 decomposition.
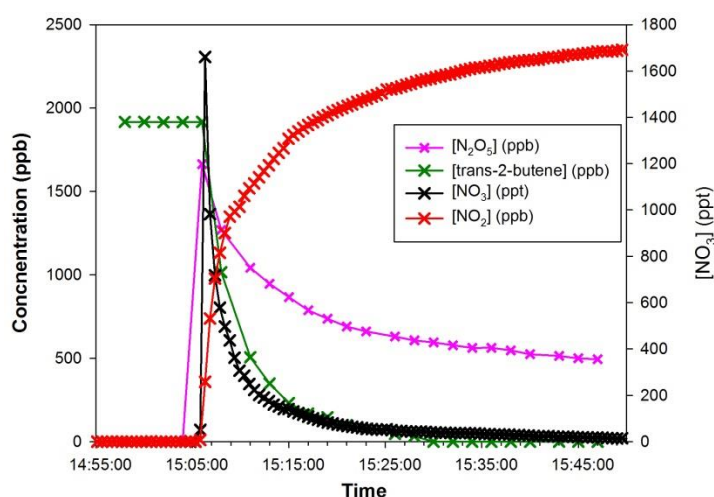

Figure 7 : Concentrations of *trans*-2-butene and NO_2 measured by FTIR (left axis) and NO_3 measured by IBBCEAS
385 **versus time (right axis) during experiment 1.**

The kinetic plot ($-\Delta$[*trans*-2-butene] vs. the product [*trans*-2-butene]$\times$[NO$_3$]$\times\Delta$t) gathering data from all experiments is presented in Fig. 8. The uncertainty on each experimental point was calculated as the sum of the relative uncertainties on [*trans*-2-butene] and [NO$_3$] for the abscissa scale (the uncertainty on the time was considered to be negligible) and as twice the uncertainty on the [*trans*-2-butene] for the ordinate scale. From this figure, it can be observed that all experiments are in good agreement. In consequence, a linear regression was performed on all data points leading to a rate constant of $(4.13 \pm 0.78) \times 10^{-13}$ cm^3 molecule^{-1} s^{-1}. The uncertainty on the rate constant was estimated as the sum of the relative uncertainties on NO$_3$ concentrations and twice the standard deviation on the linear regression. The obtained rate constant has been compared to the values from previous determinations and to the value recommended by IUPAC in Table 3.
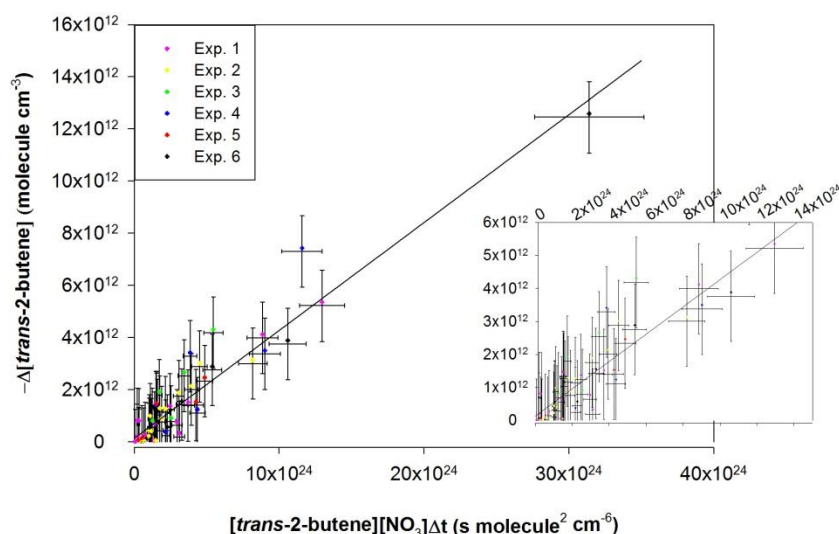
Figure 8 : Absolute kinetic plot for the reaction of *trans*-2-butene with NO$_3$ radical, showing the loss of trans-2-butene vs. [VOC]$\times$[NO$_3$]$\times\Delta$t. To the right, a zoom on the points close to the origin.

This new determination appears to be in very good agreement with the IUPAC recommendation and with previous absolute determinations, within the uncertainty. However this work does not agree with the two relative determinations which are up to 45 % lower than our determination. Nevertheless, these two values are in disagreement with all of the previous absolute determinations. These relative rate determinations are relative to the equilibrium constant K ($N_2O_5 + M \leftrightarrows NO_2 + NO_3$). A possible explanation to this disagreement would be that NO$_3$ concentration is overestimated because it was considered that the equilibrium is reached. However, the reaction with *trans*-2-butene is fast enough to significantly disturb the equilibrium and prevent it to be established. An overestimation of NO$_3$ concentrations would hence lead to an underestimation of the rate constant.

Table 3 : Comparison of the rate constant obtained for the reaction of *trans*-2-butene with NO₃ with previous determinations.

k (cm³ molecule⁻¹ s⁻¹)	T (K)	Technique*	Reference
$(4.13 \pm 0.78) \times 10^{-13}$	**293**	**(N₂O₅/CEAS)**	**This study**
$(3.90 \pm 0.78) \times 10^{-13}$ ($\Delta \log k = \pm 0.08$)	298	recommendation	IUPAC
$(3.78 \pm 0.17) \times 10^{-13}$	298	(AR/CEAS)	(Kasyutich et al., 2002)
$(3.74 \pm 0.45) \times 10^{-13}$	298	(AR/LIF)	(Berndt et al., 1998)
$(4.06 \pm 0.36) \times 10^{-13}$	298	(AR/LIF)	(Rudich et al., 1996)
$(3.88 \pm 0.30) \times 10^{-13}$	298	(AR/MS)	(Benter et al., 1992)
$(3.96 \pm 0.48) \times 10^{-13}$	298	(AR/LIF)	(Dlugokencky and Howard, 1989)
$(3.78 \pm 0.17) \times 10^{-13}$	298	(AR/LIF)	(Ravishankara and Mauldin, 1985)
$(3.09 \pm 0.27) \times 10^{-13}$	298	(RR**)	(Atkinson et al., 1984)
$(2.31 \pm 0.17) \times 10^{-13}$	300	(RR**)	(Japar and Niki, 1975)

*Indicates kinetic method (AR = absolute rate, RR = relative rate) and NO₃ measurement technique: CEAS, LIF (Laser-Induced Fluorescence) or MS (mass spectrometry) used

** Relative rate determinations are relative to the equilibrium constant $K (N_2O_5 + M \leftrightarrows NO_2 + NO_3 + M)$.

In conclusion, this agreement shows that the determination made with the IBBCEAS technique presented in this paper is correct, allowing reliable measurement of NO₃ at low concentration with good sensitivity and time resolution. This technique is now operational for application to other absolute kinetic studies.

6. Conclusions

An IBBCEAS technique has been developed and coupled to the CSA simulation chamber for the *in situ* measurement of NO₃ radicals at the ppt level. This instrument allows also monitoring NO₂ in the ppb range. Thanks to various tests, the instrument has been carefully characterized in order to identify potential bias and to define the optimal operation conditions. The performance of the instrument in terms of detection limit and uncertainties were also determined. The instrument exhibits very good detection limit for NO₃ radicals (6 ppt) for 10 seconds of integration time. This detection limit fully complies with our needs for kinetic applications.

The instrument was also validated thanks to an intercomparison experiment with the *in situ* FTIR technique. With this technique, NO₃ concentration was indirectly obtained by monitoring NO₂ and N₂O₅ concentrations and by using the well-known equilibrium constant $K (N_2O_5 + M \leftrightarrows NO_2 + NO_3)$. The concentrations measured by the two techniques were shown to be in very good agreement (better than 10%) for both NO₃ and NO₂, over a ride range of concentrations: from ppt to ppb range for NO₃ radical and from ppb to hundreds of ppb for NO₂.

Finally, this technique was used for the absolute rate determination of a well-documented reaction, *trans*-2-butene+ NO₃. The value of $(4.13 \pm 0.45) \times 10^{-13}$ cm³ molecule⁻¹ s⁻¹ found in this study is in very good agreement with the previous absolute determinations. Moreover, the good sensitivity and the good time resolution represent excellent performances allowing the use of this technique for monitoring NO₃ radicals when involved in fast reactions.

The IBBCEAS technique is now operational and will be used in further works, particularly to monitor NO₃ concentrations for absolute rate determinations of fast reactions of volatile organic compounds with NO₃ radicals.

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Answer to Anonymous Referee #1

600 First of all, the authors would like to thank the anonymous referee for this discussion and its constructive comments, corrections and suggestions that ensued. We have carefully replied to all its comments and the paper has been improved following its recommendations. All technical corrections suggested by the referee have been carefully performed. Answers have also been provided for all comments and changes have been performed accordingly. Please find below the answers to the
605 comments:

1. There are multiple missing references to figures and tables: L.104, L.150, L.171, L.196, L.213, L.227, L.247, L.295, L.336, L.349, L.356

The authors thank the anonymous referee for having pointed out these mistakes, which are probably due to an error during the submission process. All the references have been added as suggested (L.103,
610 L.152, L.173, L.202, L.220, L. 234, L.257, L.311, L. 367, L.380, L.387).

2. Further technical details should be supplied about

L.116: LED optical power

It was indeed not clearly given, but the letter K in the description of the LED indicates the bin of radiant
615 flux (K being between 390 and 430 mW).

We modified the text as follows to add this information: P.2.2 L.117: **"max. 430 mW**".

L.146: specific RH range of the experiments

Information on relative humidity has been provided P2.2 L.148: "**(RH < 1 %)**" and P.4 L.289: "**The**
620 **chamber was first filled with dry synthetic air (RH < 1%)**".

L.280-90: To form NO$_3$, either O$_3$ is needed with NO$_2$ to form NO$_3$, or N$_2$O$_5$ must be added as a precursor. This information is missing in the description. The authors should supply these experimental details along with relevant chemical equations.

625 As suggested by the referee, NO$_3$ generation protocol is now explained in the manuscript (The following text has been added between P.4 L.293 and L.298): "**NO$_3$ radicals were then formed _in situ_, using thermal dissociation of N$_2$O$_5$ (R. (1)), which was synthesized in a vacuum line following the reaction between O$_3$ and NO$_2$ (R. (2) and R. (3)). This protocol was adapted from (Atkinson et al., 1984; Schott and Davidson, 1958) and is detailed in Picquet-Varrault et al., 2009.**

630 $$N_2O_5 + M \leftrightarrows NO_3 + NO_2 + M \tag{R. 1}$$
$$O_3 + NO_2 \rightarrow NO_3 + O_2 \tag{R. 2}$$
$$NO_3 + NO_2 + M \rightarrow N_2O_5 + M \tag{R. 3}$$ "

L.270: The NO$_3$ concentration should be stated explicitly. Presumably the concentration was zero?
635 **How would other absorbing species like NO$_2$ influence the fit variability?**

It was probably not clear enough in the manuscript, but this value was estimated by considering the noise of a time profile of NO_3 concentration. The concentration was thus not zero. A precision has been added in the text (P.3.4 L.279): "**NO_3 concentration time profile**".

To answer the second question, it should be said that, in this wavelength range, spectral signatures of NO_2 and NO_3 are very different. Furthermore, NO_2 is well constrained by several thin absorption bands. The fit program handles the subtraction of NO_2 very well and the fit of NO_3 is thus not strongly impacted by its absorption. The fit variability due to NO_2 influence has been verified by constraining on the same spectrum different values of NO_2 concentration (± 100 ppb of the real value). The uncertainty on NO_2 fit can lead to an uncertainty of max. 5 ppt on NO_3 fit. This value appears to be very satisfying, considering that an error of 100 ppb on NO_2 concentrations is very unlikely.

3. It is not clear how the authors calculated certain values. These figures need to be checked:

L.177: I calculate 3.15 km for R = 0.99974 and d = 0.82 m, whereas the authors report 3.4 km. Also, it should also be clarified that this is the pathlength across the full optical cavity length (82 cm) and includes the purge volumes. When purging, the effective pathlength would be about 25% shorter.

The effective path length was calculated using Eq. 4. The value of 3.4 km is indeed an error, and it has been corrected in the manuscript (P.3.4 L.179).

This value was calculated for a full optical cavity length (82 cm) and does not include the purge volumes. A sentence has been added to specify that the purge was not used during this study (P.2.2 L.111): "**Nitrogen flush was not used in this study, but is available for further type of experiments.**"

The effective pathlength when purging is indicated in P.2.2 L.110.

L.182: I calculate 12% variation, not 20%

It was a mistake and it has been corrected in the manuscript (P.3.1 L.183).

4. The authors should comment briefly on photolysis of NO_3 by the IBBCEAS probe beam, potential surface losses of reactants, and the influence of water vapor and aerosols on the spectral analysis. These issues are likely unimportant for kinetic experiments, but should be noted for completeness.

In the used wavelength range, only one photolysis reaction can occur: for $\lambda < 710$ nm, NO_3 is dissociated into NO and O_2. Nevertheless, it has been shown that for $\lambda > 640$ nm, quantum yield of photolysis is close to 0, and thus suspected close to 0 at 662 nm. But even though NO_3 was slightly subject to photolysis locally, the volume enlightened by the probe beam is very small in comparison to the overall volume of the chamber (< 0.04%). The homogenization system allowed a mixing with the rest of the volume. Finally, the intercomparison experiment shows that this eventual loss is not significant because it would have led to lower concentrations measured by the IBBCEAS than the FTIR. So the kinetic method is not impacted by a loss of NO_3 due to photolysis. A precision has been added P.4 L.330: "**Finally, the intercomparison experiment shows that an eventual loss due to photolysis of NO_3 by the beam is not significant because it would have led to lower concentrations measured by the IBBCEAS than by the FTIR. In addition, for wavelength longer than 640 nm, which is the case here, Johnston et al., 1996 have shown that photolysis quantum yield is close to 0. It is thus expected that the photolysis of NO_3 in the used wavelength range is not occuring. Furthermore even though**

NO$_3$ was slightly subject to photolysis locally, the volume enlightened by the probe beam is very small in comparison to the overall volume of the chamber (< 0.04%) and the homogenization system allowed a mixing with the rest of the volume."

680 The authors expect that wall losses of NO$_3$ are occurring. Nevertheless, as mentioned by the referee, the absolute kinetic method consists of measuring the decay of *trans*-2-butene and to use the NO$_3$ concentration measured during the decay. Then, NO$_3$ additional losses do not affect the rate constant determination. A sentence was added in the manuscript in order to precise this point (P.5 L.360): "**It is important to notice that the absolute kinetic method used consists of measuring the decay of *trans*-**

685 **2-butene for a known concentration of NO$_3$, and not the decay of NO$_3$ radicals for a known concentration of the VOC. The method is thus not affected by NO$_3$ additional loss processes (e.g. wall losses, reactions with NO$_2$ or with peroxy radicals). Only additional losses of the VOC would lead to an overestimation of the rate constant. This was checked prior to the experiments (i.e. before in the injection of N$_2$O$_5$) and no significant loss of the VOC was observed in the timescale of the**

690 **experiment (see below).**"

There is a very strong water vapor absorption in the wavelength range used. Nevertheless, all experiments have been performed in dry conditions and no water features were present on the spectra. A sentence has been had P.2.2 L.147:"**Absorption by water vapor may be very high under atmospheric conditions.**"

695 Aerosols can also affect the IBBCEAS technique at two levels: i) in general, deposition of aerosols on the mirrors lead to a decrease of their reflectivity. That has been dealt with the nitrogen purge in case of the generation of aerosols. ii) The absorption and scattering of the beam by aerosols can lead to a major decrease of the signal in the used wavelength range. In this study, no SOA was formed in the chamber, causing no such problem. In case of the use of the technique with a system which produces

700 SOA, experimental conditions have been modified by using the nitrogen purge close to the mirrors surface and by reducing the production of SOA (by reducing the concentration of precursor).

5. It is not entirely clear how the effect of I0 stability has been determined. Is the 3% change in NO2 concentration seen at all concentrations, or is it a typical or worst case scenario? Can losses to walls

705 **influence the analysis? The residual in Fig. 3 looks very much like the NO$_2$ absorption spectrum, so why does fitting with the NO$_2$ cross-section not eliminate these features better?**

The impact of I$_0$ stability has been evaluated by performing 2 types of experiments: First, the stability of the optical system has been verified. For this purpose, long term measurements of the signal have been done, leading to intensity variations lower than 0.3% and to very small baseline distortions. Then,

710 the impact of these baseline distortions on the quantification of the absorbing species has been evaluated by measuring the evolution in time of the concentration of a stable species, here NO$_2$. A mixture of 1 ppm of NO$_2$ in synthetic air was used. A sentence has been added in the manuscript (P.3.2 L.193): "**two types of experiments have been performed: first, the stability of the optical system has been verified. For this purpose, long term measurements of the signal have been performed, leading**

715 **to variations lower than 0.3% and to very small baseline variations. Second, to verify the impact of these variations on quantification of the absorbing species, experiments were conducted […] Then, a concentration of NO$_2$ was introduced into the chamber (mixing ratios ranging between 100 ppb and 1 ppm, depending of the experiment) and the signal I(λ) was measured**" and P.3.2 L.202: "**for a concentration of 1 ppm of NO$_2$.**"

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720　It looks very unlikely that wall loss can influence the analysis, because the concentration is measured and decay would have been observed and NO_2 is very stable in the CSA chamber. In addition, NO_2 concentration is not constrained for the fit so in the case NO_2 concentration was not stable, the fitted concentration would be lower.

725　The referee is right when saying that the residual looks like NO_2 absorption spectra. It can be explained by a deformation of the baseline in time, due to the remoteness of I_0, which prevents the software to fit correctly the NO_2 spectra. It is likely that the residual is composed of both this baseline deformation and the small part of NO_2 cross section that is not correctly subtracted. The uncertainty generated by this phenomenon is said in the manuscript P.3.2 L.202 (up to 3 % for NO_2 concentration measurements).

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6. For the detection limit calculation, the authors should specify whether this calculation is for a purged or unpurged system.

The detection limit was calculated without the purge system. It has been specified P.3.3 L.214 in the manuscript: "**and an unpurged system,**"

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7. The calculation of the detection limit is based on the signal-to-noise ratio for a single wavelength. However, the spectral fit is constrained by many independent signal measurements at different wavelengths. The authors should evaluate how the multiplex nature of the measurement affects the system's detection limit.

740　We indeed used the classical method to determine the detection limit: it has been calculated on the signal-to-noise ratio for a range of wavelength between 655 nm and 666 nm, and we used the max. value of NO_3 absorption. This spectral range presents high absorption of NO_3 and weak absorption of NO_2, as said in the answer to the question 2, L.270.

It should be said that in real conditions, as shown in figure 4 in the manuscript, 6 ppt of NO_3 are very 745　close to the detection limit. The determination of the detection limit is thus coherent. Furthermore, this figure shows a spectrum with 630 ppb of NO_2, which is an important concentration. The detection limit is not degraded by the absorption of NO_2.

8. Figure 6: At low pptv, it looks like there's still an IBBCEAS signal for NO_3 but no FTIR signal. Does 750　**this divergence reflect a difference in instrument sensitivities, or equilibrium/heterogeneous chemistry, or some other cause? Please comment.**

By looking at the data, it seems that this case is for the first two points ($[NO_3]_{FTIR}$ = 3.0 ppt; $[NO_3]_{IBBCEAS}$ = 6.5 ppt and $[NO_3]_{FTIR}$ = 4.6 ppt; $[NO_3]_{IBBCEAS}$ = 6.7 ppt). For higher concentrations, the agreement is very good. It seems reasonable to think that the sensitivity of the FTIR measurement can explain this 755　difference: it is based on the measurement of NO_2 and N_2O_5. At this moment, N_2O_5 concentrations are small (approx. 8 ppb) which is close to the detection limit of the FTIR for this species (approx. 6 ppb). The uncertainty of this measurement is thus high, which can explain this difference at very small concentration of NO_3. Please note also that these concentrations are close to the detection limit of the IBBCEAS, so it will not be used in the frame of a kinetic experiment.

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9. L.313: Are spatial inhomogeneities in gas concentrations expected in the chamber?

During experiments, a homogenization system is used to prevent from inhomogeneities of the mixtures. It is constituted in three parts: (i) an injection pipe, allowing injecting all along the chamber, ii) two fans, allowing a homogenization of gas inside the chamber and iii) a close-circuit homogenization pump, which samples in one extremity to inject the mixing back in the other one. Experiments have been conducted with an inert gas (NO_2) in order to measure the mixing time. Measurements were done in several points of the chamber. This system thus allows a mixing time inferior to a minute. The good agreement between the two measurement methods in the intercomparison experiment, which measure in two different areas, shows also that there are no major inhomogeneities.

A sentence has been had to explain it better in the manuscript P.2.1 L.61: "**It is equipped with a homogenization system which is made of i) an injection pipe (4 meters long, 1 cm diameter and regularly drilled with 1 mm holes) which allows to inject the reactants all along the chamber, ii) 2 stainless steel fans allowing a homogenization of gas inside the chamber and iii) a close-circuit Teflon pump connected at both ends allowing a recirculation of the gas mixing. This system allows a mixing time below one minute.**"

10. Figure 7. NO_3 and butene losses are very rapid. Would data quality be increased by selecting reaction conditions to produce a slower reaction and more data points?

For the kinetic of *trans*-2-butene, it was not necessary to increase the frequency of the measurement because the rate constant is relatively low. For faster kinetics, experimental conditions can be changed, in particular by reducing VOC concentration or increasing the measurement frequency.

11. Most of the uncertainty in the kinetic data arises from the uncertainty in the FTIR results. Calibrated CIMS or PTR-MS might be better suited for co-measurement of VOCs. The authors may want to discuss these considerations for their kinetic studies.

We fully agree with the referee that PTR-MS measurements might have been more precise. Unfortunately, this instrument was not available in our group when we performed the experiments.

Minor corrections:

L.27: either "the NO3 radical has. . ." or "NO3 has. . ."

It has been done.

L.38: "one of the reasons for this. . ."

It has been done.

L.46: "progress has been made"

It has been done.

L.53: "For this purpose,"

It has been done.

L.63: "presented in detail"

It has been done.

800 **L.76: "NO3 was"**

It has been done.

L.101: "planar/concave"

It has been done.

L.106: "prevent adsorption of. . ."

805 It has been done.

L.115: Caption Caption should note that collimating lenses and curved mirrors are not shown on the beam injection side.

It has been done.

L.127: "LED current is fixed at 900 mA"

810 It has been done.

L.132: "to focus the beam"

It has been done.

L.143: hyphen in reference

It has been done.

815 **L.175: "justifies...wide scale". It is unclear what is meant by this. That it should be measured over a wide wavelength range, or that it is not necessary to do so?**

The sentence has been changed to: "which justifies that it is necessary to measure it on a wide wavelengths range".

L.182: "prior to each experiment"

820 It has been done.

L.196: "up to 3 % in NO2 concentration"

It has been done.

Figure 3. Change "residue" to "residual" in figure titles and legends.

It has been done.

825 **L.280: "was first filled with dry synthetic air"**

It has been done.

L.281: "Air Liquide NO2,"

N20 (should be read N-20) is actually the name of the cylinder.

L.287: NO2 and N2O5 concentrations

830 It has been done.

L.300: in IBI and in the spectral treatment

It has been done.

L389: performance

It has been done.

835 **L.396: Clarify "from ppt to ppt range for NO3"**

It has been done.

L.400: Monitoring NO3 radicals

It has been done.

L.409: intercomparison

840 It has been done.

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Answer to Anonymous Referee #2

First of all, the authors would like to thank the anonymous referee for this discussion and its constructive comments, corrections and suggestions that ensued. We have carefully replied to all its comments and the paper has been improved following its recommendations. All technical corrections suggested by the referee have been carefully performed. Answers have also been provided for all comments and changes have been performed accordingly. Please find below the answers to the comments:

1. The kinetics analysis ignores any loss of NO3 other than bimolecular reaction of NO_3 with Alkene. This is most probably an over simplification. For example, what about wall loss, there could also be diffusion outside of the analysis region. Using a simple model, the addition of a first order loss of NO_3, still results in a straight line from a plot of the change in alkene vs the product of [alkene].[NO₃] delta t (as in figure 8), however the retrieved rate coefficient is lowered and this impact gets worse as the wall loss increases. So if there is wall loss the retrieved rate coefficient does not equate to the bimolecular rate coefficient. How have the authors taken into account wall loss or any other first order loss process? Have they modeled their system? Can they show what the quantitative impact on the retrieved rate coefficient will be? Have they made any attempt to evaluate first order losses of NO_3?

We may have not explained clearly enough the kinetic method, because it appears that the referee did not understand it. NO_3 is indeed subject to losses other than the reaction with the VOC, for example wall losses in the chamber, reaction with NO_2, with peroxy radicals… Nevertheless, this method consists of monitoring the decay of the *trans*-2-butene and not the one of NO_3, and this decay is plotted as a function of $[VOC] \times [NO_3] \times \Delta t$. Here, $[NO_3]$ is measured with the IBBCEAS technique as a function of the time (and not modelled), so we don't care about additional loss processes. What is important here is the measured concentration of NO_3 radicals by the time the consumption of the VOC is measured. So, NO_3 can be consumed by other processes without affecting the result. Bu if the VOC is consumed by other loss processes (e.g. wall losses), this could affect the results. For this reason, *trans*-2-butene stability has been checked before every oxidation, and no significant loss beside dilution was observed. It allowed the authors making the hypothesis that *trans*-2-butene was only consumed by reaction with NO_3.

2. The authors claim that there is good agreement with literature, however, their rate coefficient is the fastest that has been reported in the literature. Also, given that this rate coefficient is probably not simply the rate coefficient for the bimolecular process, most probably a lower limit as result of other first order loss proves, the true bimolecular rate coefficient in their system is likely to be even faster. Faster rate coefficients normally worry kineticists, if one looks at the absolute rates reported in table 3 the unweighted average is 3.87×10^{-11} with a standard deviation of 0.12×10^{-11} i.e. statistically the reported rate coefficient is higher than the current experimental database. Can the authors explain why this is the case?

We are a bit surprised by this comment. Isn't (3.87 ± 0.12) in agreement with (4.13 ± 0.78)? Agreement between values has to be considered by taking into account the uncertainties, otherwise it has no sense. And by taking them into account, our determination is in agreement with the unweighted average of absolute determinations and with the IUPAC recommended value. But it is also in agreement with all individual absolute determinations. The only disagreement which has been noticed

is with relative determinations which are significantly lower than other determinations and where not taken into account for the IUPAC recommended value. In conclusion the authors affirm that our determination is in good agreement with literature.

3. In Table 1, Deff is given as 82 which is simple the geometric distance, if I am not mistaken. Will that be the case? A purge flow is used to protect the mirrors and this will impact the effective path length, this is also likely to be a function of pressure and flow of purge gas. Have experiments been performed to quantify the Deff as a function of purge gas flow rate and pressure?

All experiments presented here were performed without a purge flow so the effective distance was considered to be the geometric distance (82 cm). A sentence clarifying this statement has been added in the manuscript (P.2.2 L.111): "**Nitrogen flush was not used in this study, but is available for further type of experiments.** "

Nevertheless, experiments have been conducted to quantify the impact of purge gas flow and pressure:

- Due to nitrogen purge, pressure is only slightly rising: the purge flow is 300 mL min^{-1}, and 300 mL represents a variation of 0.03 % of the total volume of the chamber. This flow is compensated by the sampling of measurement instruments. No pressure variation was ever measured due to the purge gas flow. Nevertheless, characterization experiments using a known NO_2 concentration were conducted, in order to quantify the impact of major pressure variation on the pathlength. They showed no significant impacts of pressure on the pathlentgth. Pressure variations of 10 mbar in the chamber lead indeed to decrease of the path length of less than 1 %.
- Using also a known concentration of NO_2, experiments have been performed by adding different flows of purge, in order to determine an efficient flow, i.e. which does not induces a major decrease of the pathlength and a progressive dilution in the measurement area. They showed that the optimal flow is 300 mL min^{-1}, because it efficiently protects the mirrors and induce no long term dilution. With this flow, Deff is decreasing to 62 cm (as it is written in the manuscript L.110), adding an uncertainty of 5 % to the nitrate radical concentration. A sentence to explain these experiments has been added in the manuscript P.2.2 L.108: "**This flow rate has been optimized in order to efficiently protect the mirrors while limiting the dilution of the mixture in the measurement area.**"

4. On line 290 the authors state "these parameters were therefore precisely monitored during the experiment leading to the value of 2.17 × 10-11 cm3.molecule-1 at 298K and at 1030 mbar." Is that the value of the equilibrium constant? If so, what are the errors? They need to be stated. Can the authors show that within experimental error that agrees with those reported in Atkinson et al., 2004, indeed how does it compare to the recommended IUPAC /JPL recommended values?

The value of 2.17 × 10^{-11} cm^3.molecule^{-1} is indeed the equilibrium constant. This value was not measured in our study, but calculated using the parametrization recommended by the IUPAC (Atkinson et al., 2004) at a pressure and temperature which were measured during the experiments. So, here, it appears that there is a misunderstanding. By "parameters", we mean temperature and pressure and not the rate constant. In order to make it clearer, the sentence has been modified P.4 L. 305 in the manuscript: "**These two parameters were therefore precisely monitored during the experiment**

allowing calculating an equilibrium constant of 2.17 × 10^{-11} cm^3.molecule^{-1} at 298K and at 1030 mbar, using IUPAC database parameters (Atkinson et al., 2004)."

5. In Figure 6, the authors report linear regression between the FTIR and the BBCEAS. Can they provide errors on the slopes? Also include a description of those errors, e.g. are they 1 sigma just from the linear fit, or do they take into other experimental errors?

Errors on the slopes are now provided. The values were calculated to be 1.0 ± 0.2 for NO$_2$ and 1.1 ± 0.3 for NO$_3$ by considering the statistical error on the slope, which is twice the standard deviation, and the sum of the systematic relative errors on FTIR and IBBCEAS measurements.

To explain it, sentences have been written in P.4 L.313: "**Here, the overall uncertainty was calculated as the sum of the statistical error on the slope (twice the standard deviation, 4 %) and systematic errors on FTIR (i.e. on IBI$_{NO2}$, 4 %) and IBBCEAS measurements (which includes uncertainties on NO$_2$ cross sections and on the mirrors reflectivity, 9%).**"and in P.4 L.320 "**The error is calculated with the same method as for NO$_2$.**"

6. Finally what are the absolute errors on the rate coefficient? The authors report a simple error analysis based on the line of best fit, i.e. what is the total error? The authors need to take into account errors in flows, absorption cross section etc etc

The authors have considered that the statistical error, calculated as twice the standard deviation on the linear regression and which takes into account the dispersion of the experimental points, is mainly due to spectra treatment uncertainties. However, we agree that this statistical error does not include systematic ones, such as errors on cross sections. So we have added the uncertainty on NO$_3$ concentration. As described in the manuscript, this uncertainty was estimated to be 9% and corresponds to the sum of NO$_2$ and NO$_3$ cross sections errors and the uncertainty of NO$_2$ concentration for reflectivity measurement. By summing the statistical error (10%) and the uncertainty on NO$_3$ concentration (9%), we obtain an overall uncertainty of 0.78 × 10^{-13} cm^3 molecule^{-1} s^{-1}. This has been corrected in the manuscript L.391 and in Table 3. It must be noticed that because *trans*-2-butene concentration is both on x and y axes, the error on IBI is not considered here.

An explanation has been added P.5 L.392: "**The uncertainty on the rate constant was estimated as the sum of the relative uncertainties on NO$_3$ concentrations and twice the standard deviation on the linear regression.**"