A compact, high-purity source of HONO validated by Fourier Transform Infrared and Thermal Dissociation Cavity Ring-down Spectroscopy

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Abstract

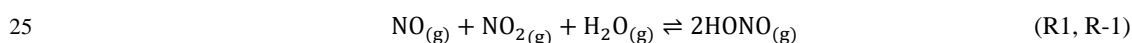
A well-characterized source of nitrous acid vapour (HONO) is essential for accurate ambient air measurements by instruments requiring external calibration. In this work, a compact HONO source is described in which gas streams containing dilute concentrations of HONO are generated by flowing

10 hydrochloric acid (HCl) vapour emanating from a permeation tube over continuously agitated dry sodium nitrite ($NaNO_2$) heated to 50 °C. Mixing ratios of HONO and potential by-products including NO, NO_2 and nitrosyl chloride (ClNO) were quantified by Fourier Transform Infrared (FTIR) and thermal dissociation cavity ring-down spectroscopy (TD-CRDS). A key parameter is the concentration of HCl, which needs to be kept small (< 4 ppmv) to avoid ClNO formation. The source produces gas streams containing HONO in

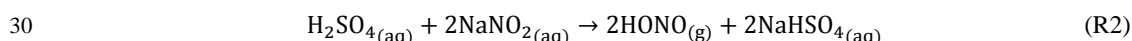
15 air in > 97% purity relative to other nitrogen oxides. The source output is rapidly tuneable and stabilizes within 90 min. Combined with its small size and portability this source is highly suitable for calibration of HONO instruments in the field.

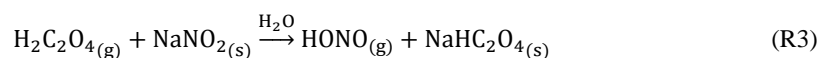
1 Introduction

20 The generation of nitrous acid vapour (HONO$_{(g)}$) free of other nitrogen oxides has been a long-standing
challenge to atmospheric chemists (Table 1). Such sources are needed in kinetic studies, for absorption
cross-section measurements, and for the calibration of field instruments. Nitrous acid is not suitable for
permeation devices since HONO is difficult to prepare in low concentration and in high purity and dis-
proportionates to nitric oxide and nitrogen dioxide via equilibrium (1).

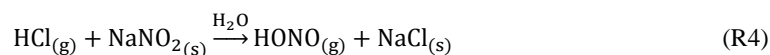
25
$$NO_{(g)} + NO_{2\,(g)} + H_2O_{(g)} \rightleftharpoons 2HONO_{(g)} \qquad\qquad (R1, R-1)$$

In fact, early experiments relied on the NO, NO$_2$ and H$_2$O vapour equilibrium to form HONO in situ (King
and Moule, 1962; Cox, 1974; Stockwell and Calvert, 1978).

Cox and Derwent prepared HONO vapours in ~50% purity by flowing nitrogen over an aqueous solution
containing 0.1 M sodium nitrite with and 1.4%-2.5% sulfuric acid (Cox and Derwent, 1976).

30
$$H_2SO_{4\,(aq)} + 2NaNO_{2\,(aq)} \rightarrow 2HONO_{(g)} + 2NaHSO_{4\,(aq)} \qquad\qquad (R2)$$

Braman and de la Cantera sublimed oxalic acid onto solid sodium nitrite and were able to produce gas
streams in >50% and up to 90% purity as long as water was present (Braman and De la Cantera, 1986).

$$H_2C_2O_{4\,(g)} + NaNO_{2\,(s)} \xrightarrow{H_2O} HONO_{(g)} + NaHC_2O_{4\,(s)} \qquad\qquad (R3)$$

Subsequent refinements included a flow reactor design in solutions of sulfuric acid and sodium nitrite were
35 dynamically mixed, producing a stable output in >90% purity (Taira and Kanda, 1990).

Febo et al. (1995) developed a method in which gas-phase hydrochloric acid emitted from a permeation
device was quantitatively reacted in a humidified gas stream with solid sodium nitrite, producing a stable
and highly pure (>99.5%) HONO output.

$$HCl_{(g)} + NaNO_{2\,(s)} \xrightarrow{H_2O} HONO_{(g)} + NaCl_{(s)} \qquad\qquad (R4)$$

40 The output of sources based on R4 has been characterized by differential optical absorption spectroscopy
(DOAS) (Febo et al., 1995; Stutz et al., 2000), Fourier transform infrared (FTIR) spectroscopy (Brust et al.,
2000; Schiller et al., 2001), tuneable diode laser absorption spectroscopy (TLDAS) (Schiller et al., 2001),
incoherent cavity-enhanced absorption spectroscopy (IBBCEAS) (Roberts et al., 2010), and long-path
optical absorption photometry (LOPAP) (Ren et al., 2010). Stutz et al. (2000) noted that concentrations of

45   HONO should be kept low ($< 10^{14}$ molecules cm^{-3}; ~4 ppmv) to avoid disproportion of HONO via R-1.

Perez et al. (2007) examined a source based on the design by Febo et al. (1995) using thermal-dissociation chemiluminescence (TD-CL) detection of NO and reported that their source co-emitted nitrosyl chloride (ClNO); they opted to generate HONO from reaction of H_2SO_4 aerosol with $NaNO_2$(s) instead. In contrast,

50   Roberts et al. (2010) observed little or no production of ClNO (<4.5%) and quantitative conversion of HCl to HONO, at odds with Perez et al. (2007) but in agreement with the original work by Febo et al. (1995). As far as we know, this inconsistency is unexplained to date.

As part of our development of an IBBCEAS for HONO measurement, our lab recently constructed a HONO source using an aqueous solution of $NaNO_2$ buffered with oxalate to pH 3.74 (Jordan and Osthoff, 2020).

55   This source generated HONO in trace amounts suitable to our needs but co-emitted NO and NO_2 in similar ratios as reported by Braman and De la Cantera (1986). Furthermore, this type of source could not be turned off (i.e., had to remain under continuous N_2 flow) and required long times to stabilize, motivating us to develop an alternative HONO generation method. Ambient air HONO measurement techniques that need to be externally calibrated often poorly agree with each other (Crilley et al., 2019), further motivating the

60   development of a well-characterized, compact and portable HONO source and to improve the understanding of conditions needed to suppress formation of undesired side products such as ClNO.

In this work, we describe the implementation of a compact, high-purity HONO generation device based on R4 for field deployment and laboratory experiments and characterized its output by FTIR and thermal dissociation cavity ring-down spectroscopy (TD-CRDS). Conditions to avoid generation of impurities such

65   as ClNO are identified. The performance of this source in comparison to existing methods is discussed.

2 Materials and methods

2.1. Generation of gas streams containing HONO

A practical challenge is the generation of stable and dilute gas streams containing HCl. Calibration gas

70   cylinders are the most straightforward method to use for this purpose but may require periods of up to 10 days to stabilize (Roberts et al., 2010). Furthermore, low-concentration HCl gas cylinders are expensive to source, such that we chose to generate HCl gas through the use of relatively inexpensive permeation devices.

In the setup by Febo et al. (1995), which was also implemented by Brust et al. (2000), Stutz et al. (2000) and Ren et al. (2010), a gas stream of N_2 is passed through Teflon tubing immersed in an HCl bath. This setup involved a rather large 1 L vessel containing liquid HCl which is somewhat impractical and potentially hazardous in a field setting. With portability in mind, we decided to construct a permeation tube containing HCl(l). Literature does not inform as to suitable tube dimensions, which were determined through trial-and-error.

Three HONO sources were constructed and evaluated. In the first ("source 1"), a 24 cm long, polytetrafluoroethylene (PTFE) permeation tube (Chromatographic Specialties C111LE; wall thickness 1.5 mm; outer diameter (o.d.) 7.8 mm) was filled with 2.5 mL of 37% HCl (Sigma-Aldrich) and sealed at both ends with PTFE plugs held in place by stainless steel compression rings. The permeation tube was placed inside a glass chamber (VICI Dynacalibrator Model 120) whose temperature was controlled at 25.0 °C and which was continuously flushed at a flow rate of ~0.15 L min⁻¹ with room air scrubbed using activated charcoal for a period of several days. The use of scrubbed air ensured that the gas stream contained water to maintain efficient HONO production (Schiller et al., 2001).

The HCl output was diluted in a gas stream of O_2 (~20 mL min⁻¹). Roughly 1.5 g of solid $NaNO_2$ was placed inside a 50 cm long loop of 3/16" (0.476 cm) inner diameter (i.d.) and ¼" (0.635 cm) o.d. fluorinated ethylene propylene (FEP) Teflon tubing downstream from the HCl addition point. This section could be manually bypassed with a pair of 3-way valves (Entegris) to turn HONO production on or off. A Teflon filter (Pall, 2 μm pore size and 47 mm diameter) inside a Teflon filter holder (Cole-Parmer) was placed downstream of the $NaNO_2$ powder.

In the second source ("source 2"), the $NaNO_2$ was placed in a two-neck, 50 mL round bottom flask which was covered in aluminium foil to prevent photolysis of nitrite and nitrous acid and was externally heated to a temperature of 50 °C using a water bath and mechanically agitated using a magnetic stir bar as described by Febo et al. (1995).

The third source ("optimized source") used the same setup as source 2, except that a shorter permeation tube (length 6 cm) containing 0.35 mL of HCl was used. This setup is depicted in Figure 1. The overall dimensions are 28 cm × 28 cm × 36 cm and its modular design ensures easy transport to and from the field.

To deliver HONO in atmospheric concentrations (i.e., < 10 ppbv), a portion of the source output was directed towards waste with the aid of a pump and a needle valve. The output concentration could be rapidly changed by adjusting the position of the needle valve. The remaining output was diluted using scrubbed air and directed towards the instruments.

2.2 Analysis of HONO source output by FTIR

Gas streams exiting the HONO sources were analysed using an FTIR spectrometer (Bruker Tensor 27) equipped with a liquid nitrogen cooled mercury cadmium telluride (MCT) detector and a White multi-pass gas cell with a 6.4 m optical path length and an internal volume of 0.75 L (Gemini Scientific Instruments, Venus series) in a similar fashion as described earlier (Taha et al., 2013). Spectra were acquired continuously at a time resolution of 30 s. Background spectra were recorded with HCl inline but with the $NaNO_2$ bypassed, which resulted in the observed spectra showing the change in trace gas concentrations.

Mixing ratios of trace gases were determined from fits (by least squares error minimization over a selected wavelength range) of reference spectra from the Pacific Northwest National Laboratory (Sharpe et al., 2004), multiplied by the respective mixing ratios as variables, to the observed spectra. The reference spectra are provided in units of ppmv at atmospheric pressure per meter of optical length, necessitating a correction factor of 6.4 accounting for the actual path length and a pressure correction factor, for which the pressure inside the multi-pass cell was monitored using a pressure transducer (Omegadyne PX419–015A5V).

The FTIR limits of detection (LODs) are specific to each molecule (due to differing absorption cross-sections) and are in the 100 to 300 parts-per-billion (by volume; ppbv; 10^{-9}) range. For example, the 1σ precision of NO_2 data was ±100 ppbv, yielding a 3σ LOD of 300 ppbv.

The spectral resolution of the FTIR was ~0.5 cm^{-1}, which does not suffice to fully resolve the absorption lines of HCl, H_2O and NO; their FTIR derived mixing ratios are hence lower limits.

2.3 Analysis of HONO source output by TD-CRDS

The (diluted) HONO source output was also analysed using a four-channel TD-CRDS instrument (Odame-Ankrah, 2015). Briefly, mixing ratios of NO_2 are quantified by absorption at 405 nm. Ring-down time constants in the absence (τ_0) and presence (τ) of NO_2 are converted to concentrations (N) using equation (2) where c is the speed of light, σ is the NO_2 absorption cross-section, and R_L is a correction factor (~1.2) accounting for mirror purge flows (Paul and Osthoff, 2010).

$$N = \frac{R_L}{c\sigma}\left(\frac{1}{\tau} - \frac{1}{\tau_0}\right) \tag{1}$$

Concentrations are converted to mixing ratios using the ideal gas law. Other NO_y components are converted to NO_2 in separate channels and are quantified by difference.

For this study, the instrument was operated as follows: Mixing ratios of NO_2 were monitored using a room temperature, 1/4" (0.64 cm) o.d. and 1/8" (0.32 cm) i.d. FEP Teflon inlet. Mixing ratios of NO_x were quantified on a second channel with a room temperature Teflon inlet and by adding O_3 (mixing ratio after addition ~3 ppmv) to titrate NO to NO_2 (Odame-Ankrah, 2015; Fuchs et al., 2009). A third channel was operated with a 1/4" (0.64 cm) o.d. quartz inlet heated to 600 °C. Ozone was added between the quartz inlet and the CRDS cell to quantify NO_y, including HONO (Jordan and Osthoff, 2020; Wild et al., 2014; Perez et al., 2007). The fourth inlet was operated at 350 °C and with added O_3. This channel quantified $NO_x + \Sigma PAN + \Sigma AN + ClNO + ClNO_2$ and was used as a HONO reference channel. Each channel sampled at a flow rate of ~0.8 slpm through ~5 cm short, 1/16" i.d. Teflon flow restrictors placed inline after the heated quartz sections (if existing) and before a 50 mm Teflon (Pall Teflo, 2 μm pore size) housed in a Teflon filter holder (Cole Parmer). The laser pulse repetition rate was 1500 Hz, and 1500 ring-down events were averaged to produce 1 s data.

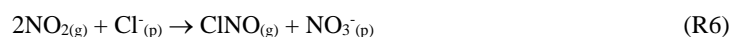
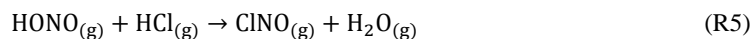
3 Results

3.1 Development of a compact, high-purity HONO source

The first design tested (source 1) produced an output of ~38 ppmv of HONO from a similar amount (>38 ppmv) of HCl (Figure S1A). This HONO mixing ratio is a factor of ~10 larger than the maximum recommended by Stutz et al. (2000) to avoid partial conversion of HONO to NO_2 and NO via R5. Consistently, the source output contained ~4.0 ppmv of NO_2 (Figure S1A) and an analogous amount of NO (Figure S1B). Furthermore, the spectrum contained an absorption feature around 1808 cm^{-1}, reproduced by ~15.5 ppmv of ClNO.

Nitrosyl chloride can be produced by reaction of HONO with gas-phase or wall-adsorbed HCl (Zhang et al., 1996; Wingen et al., 2000) or by reaction of two equivalents of NO_2 with particulate phase Cl$^-$ (Weis and Ewing, 1999), though the relatively low mixing ratios of NO_2 in this system makes this reaction less likely.

$$HONO_{(g)} + HCl_{(g)} \rightarrow ClNO_{(g)} + H_2O_{(g)} \qquad (R5)$$

$$2NO_{2(g)} + Cl^-_{(p)} \rightarrow ClNO_{(g)} + NO_3^-_{(p)} \qquad (R6)$$

The FTIR spectrum contained ~50 ppmv more H_2O than the reference spectrum, i.e., more moisture than emitted by the permeation tube with $NaNO_2$ bypassed. While this is consistent with ClNO production via

R6, the H$_2$O data are likely not meaningful since the H$_2$O mixing ratios changed slowly over time (data not shown), possibly because of slow equilibration with the inner walls of the tubing and multi-pass cell and perhaps also because of water's presence in the FTIR optical path outside the multi-pass cell at a concentration that may have drifted.

165 Febo et al. (1995) reported lower production of NO$_x$ by mechanically agitating the NaNO$_2$ using a stir bar (to break up pockets of high [HONO]) and by heating the reaction vessel to 50 °C; these steps were incorporated in the second design (source 2). The FTIR spectrum of its output showed a consumption of > ~42 ppmv HCl and production of ~48 ppmv HONO, ~8.0 ppmv of NO$_2$, ~6.5 ppmv of ClNO, ~8 ppmv of NO, and ~50 ppmv of H$_2$O (Figure S2), an improvement over source 1 but still inadequate.

170 A shorter HCl permeation tube was used in the third design, with the expectation that the lower HCl concentrations results in reduced generation of side products via R5 and R6. A sample FTIR spectrum of this source's output is shown in Figure 2. With a freshly prepared HCl permeation tube, > ~2.5 ppmv HCl were consumed and ~3.0 ppmv of HONO were produced. The mixing ratios of undesired side products (i.e., NO, NO$_2$, ClNO, and HNO$_3$) were below their respective FTIR detection limits (Figures 2, S3 and

175 S4).

3.2 Analysis of source output by TD-CRDS

To better constrain the mole fractions of impurities, the source output was analysed by TD-CRDS. Figure 3 shows an example experiment. The TD-CRDS sampled scrubbed air before 21:38 and after 21:58 to determine the ring-down time constants in the absence of absorbers, τ_0. In the time periods in between,

180 varying amounts of the HONO output were sampled. There was little response other than in the NO$_y$ channel, the only channel sensitive to HONO. Scatter plots of NO$_2$, NO$_x$ and NO$_x$ + ΣPAN + ΣAN + ClNO + ClNO$_2$ ("HONO ref") against NO$_y$ (Figure 3, inserts) have slopes of (1.29±0.06)%, (1.54±0.06)% and (2.28±0.04)%, respectively, from which a source purity of >97.3% was deduced.

185 When concentrations changed, the 90%-10% rise (and fall) times of the TD-CRDS was < 3 s, a time constant identical to changes in NO$_x$ concentration and indicating a "well-behaved" inlet, i.e., the absence of inlet memory effects and fast equilibration with the inner walls of the inlet. The fast response contrasts with the rise (and fall) times of HNO$_3$ of > 180 s (data not shown). Furthermore, the response in a channel sampling from an inlet heated to 600 °C without added O$_3$ was the same as the that of the NO$_2$ channel

190 (data not shown), consistent with the absence of HNO$_3$ observed by FTIR spectroscopy (Figure S4).

A quartz inlet temperature scan when the TD-CRDS was sampling a constant concentration of HONO is shown in Figure 4. The experimental TD curve was reproduced by a fit to equation (2), which is based on equation (4) by Paul et al. (2009).

$$[\text{Nardqudc}]_{\text{sns'k}} < [NO_x] + [ClNO]\left(0, d^{\cdot @ClNO » d^{\frac{-E_{A,ClNO}}{RT} » tres}}\right) + [HONO]\left(0, d^{\cdot @HONO » d^{\frac{-E_{A,HONO}}{RT} » tres}}\right) \quad (2)$$

195 Here, A_{HONO} and $E_{A,HONO}$ are Arrhenius parameters for the TD of HONO taken from Tsang (1991), A_{ClNO} and $E_{A,ClNO}$ are from Baulch et al. (1981), and t_{res} is the residence time of the gas in the converter at the converter set temperature T. A fit of equation (2) to the observation (superimposed as a black line in Figure 4) using the software package Igor Pro (Wavemetrics) gave $t_{res} = (77.6\pm0.1)$ ms, $[NO_x] = (0.03\pm0.30)$ ppbv, $[ClNO] = (0.27\pm0.34)$ ppbv, and $[HONO] = (49.97\pm0.18)$ ppbv, suggesting a source purity of $(99.4\pm0.9)\%$.

200 In our experience, the TD inflection points vary between TD-CRDS channels (and depend on sample flow rates), captured by the t_{res} parameter that is specific to each channel. For the data shown in Figure 4, HONO fully dissociates at a temperature of ~570 °C, and the inflection point is at ~500 °C. Nitrosyl chloride is predicted to fully dissociate at ~330 °C (inflection point at ~288 °C), such that the temperature of 350 °C chosen for the HONO ref channel (Figure 3) is justified. The blue, dashed line superimposed in Figure 4

205 shows a simulation of 5 ppbv ClNO and 45 ppbv HONO to demonstrate the relative positions of the TD of ClNO and HONO. We attempted to verify the ClNO TD profile experimentally, but the unoptimized source's ClNO outputs were not sufficiently stable, and those attempts were abandoned. Regardless, it is clear from Figure 4 that the contribution of ClNO to the total source output is negligible.

As can be seen from the insert in Figure 4, HONO contributes a tiny amount to the TD-CRDS signal at
210 350 °C (~4 pptv in the simulation of 50 ppbv HONO). This contribution may differ between quartz heaters and channels and could have impacted by TD-CRDS using multiple parallel channels presented at the beginning of this section and in Figure 3, leading to an overestimation of the total amount of impurities present in that experiment. The temperature scan data shown in Figure 4, on the other hand, were collected using a single quartz heater and are thus more accurate.

215 **3.3 Source stability**

The source output gradually decreased over a time scale of weeks of continuous use, which was rationalized by the visible depletion of the HCl permeation tube. However, the source output remained stable and reproducible on shorter time scales. An example time series is shown in Figure 5, which was acquired after

1 month of intermittent use. After a 1.5 hours stabilization period, the source produced a stable output of

220   1.57 ppmv of HONO (from >1.0 ppmv of HCl) with a precision of ±35 ppbv.

4. Discussion

In this work, a small footprint, portable, stable, rapidly tuneable, and high-purity (>97%) HONO source has been described. The source achieves compactness by using a diffusion tube containing 0.35 mL

225   concentrated HCl rather than a 1 L HCl bath. The latter is a considerable hazard since an HCl spill can cause severe burns of skin or permanent eye damage, and hydrochloric acid vapours are toxic to inhale. Hence, smaller volumes are preferred when deploying to the field. The source stabilizes within a time frame of 1.5 h (Figure 5) which is fast compared to, for example, the 10 day stabilization period reported by Roberts et al. (2010). The source output is also rapidly tuneable (Figure 3) which allows expedient

230   generation of calibration plots. These properties make it a useful setup for in-field instrument calibrations.

The combination of FTIR and TD-CRDS provided a unique and powerful tool set to analyse the purity of the source's output. This work has shown that HONO sources based on R4 can generate NO, NO_2 and ClNO as by-products via R-1 and R5. Febo et al. (1995) and Roberts et al. (2010) have presented HONO sources based on R4 and demonstrated stoichiometric conversion of HCl to HONO. The results in this work suggest

235   that devices relying solely on loss of HCl for calibration of HONO output should be used with caution as co-production of NO, NO_2, and ClNO is a real possibility. In our setup, formation of ClNO could be avoided by keeping the HCl concentration low (i.e., < 4 ppmv), which puts an important constraint on sources generating HONO by R4. With this in mind, it is instructive to scrutinize the earlier work in regard to side product formation. It has generally been noted that the presence of moisture is needed to ensure a high

240   HONO output. We suggest that water is needed in part because ClNO is prone to hydrolysis (Karlsson and Ljungstrom, 1996) such that the product distribution is shifted towards HONO.

$$ClNO_{(g)} + H_2O_{(g)} \rightarrow HONO_{(g)} + HCl_{(g)} \tag{R-5}$$

Roberts et al. (2010) observed no ClNO production, likely because they kept their concentrations relatively low (<900 ppbv) and humidified their gas streams, though they appear to be just below the threshold of

245   possible ClNO production. Perez et al. (2007), on the other hand, observed ClNO but did not disclose details as to how they operated their source, i.e., if they worked under trace conditions and with humidified gases, though it is likely that either or both conditions were not met. Schiller et al. (2001) bubbled N_2 through liquid HCl and would have certainly achieved high enough HCl concentration to generate ClNO with their

setup. However, they reported achieving higher HONO output in the presence of moisture. Hydrolysis of

250 ClNO (R-5) may be a (partial) explanation why the presence of H_2O improved the output of their HONO.

The present HONO source has a large dynamic range (0 - 1 ppmv), well above the HONO concentrations found in ambient air. For field calibrations in future, it may be desirable to construct a permeation tube with lower HCl output, for example by lowering the HCl concentration in the permeation tube.

Data availability

255 The data used in this study are available from the corresponding author upon request (hosthoff@ucalgary.ca).

Author contributions

NJG and HDO designed the experiments and carried them out.

Competing interests

260 The authors declare that they have no conflict of interest.

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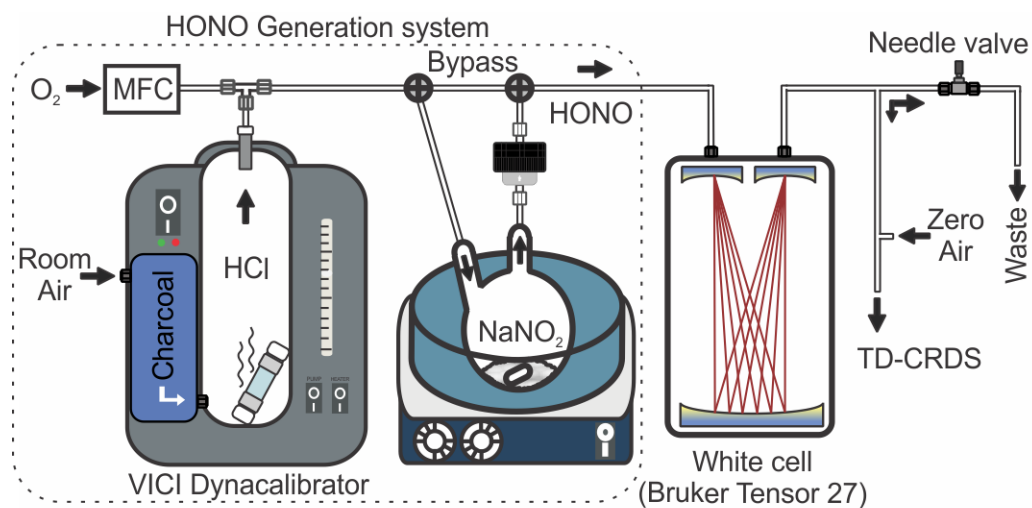

350 **Figure 1.** Schematics of the experimental setup. The high-purity HONO generation system is shown on the left (dotted border). MFC = mass flow controller.

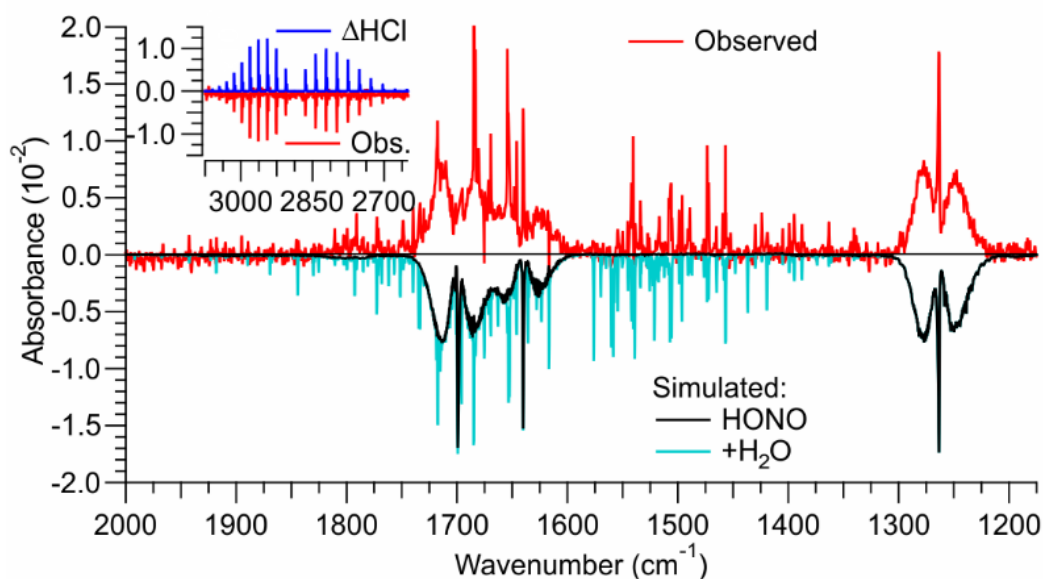



Figure 2. Infrared spectrum (shown in red colour) of a gas stream containing HONO generated by reaction

355   of $HCl_{(g)}$ with $NaNO_{2(s)}$ after the source was optimized. The reference spectrum was collected when $NaNO_2$ was bypassed, i.e., contained HCl. Literature spectra (Sharpe et al., 2004) were multiplied by the optical path length of 6.4 m and mixing ratios of the identified trace gases until they reproduced the observed spectrum. The optimized source delivered 3.0 ppmv of HONO from >2.5 ppmv of HCl. The spectrum also contained ~4.0 ppmv of H_2O. The HCl and H_2O concentrations are underestimates of their true

360   concentrations since their absorption lines are narrower than the resolution of the FTIR of 0.5 cm^{-1}.

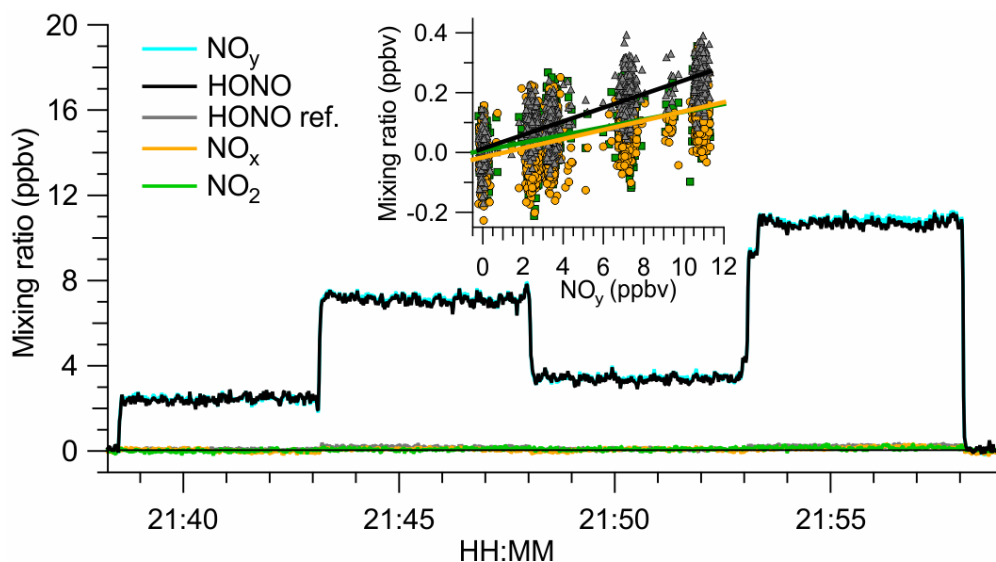



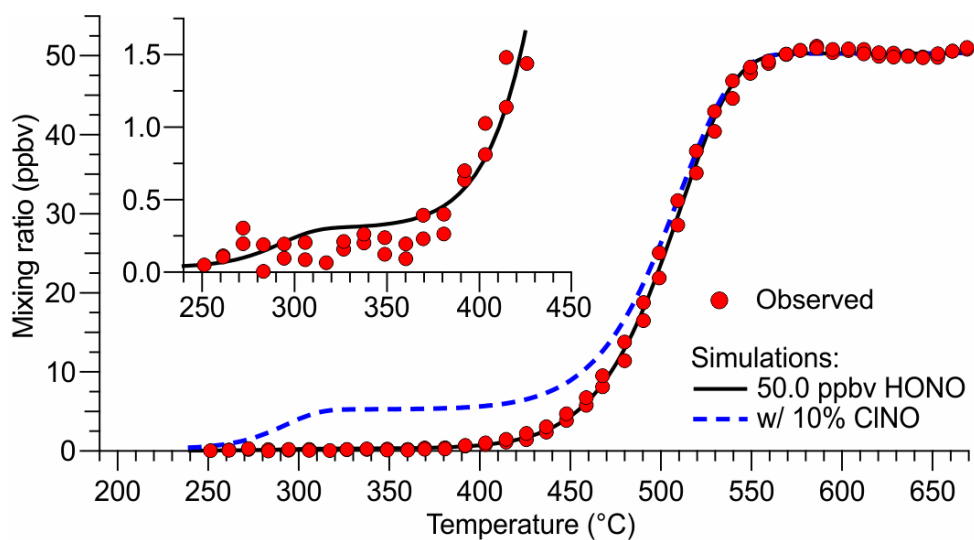
Figure 3. Analysis of the HONO source output by TD-CRDS. The TD-CRDS sampled scrubbed "zero" air before 21:38 and after 21:58. The HONO source output was varied by incrementally opening (or closing) the bypass valve. The HONO mixing ratio was calculated by subtracting the response of the "HONO reference" from the NO_y channel. The insert shows scatter plots of NO_2, NO_x and "HONO ref" against NO_y. Slopes of $(1.29\pm0.06)\%$ for NO_2 (points shown in green), $(1.54\pm0.06)\%$ for NO_x (data points shown in orange) and $(2.28\pm0.04)\%$ for "HONO ref") (points shown in grey) were determined, respectively.

370

Figure 4. Inlet temperature scan when the TD-CRDS was sampling a constant concentration of HONO (data points shown in red). The TD profile was reproduced by a fit to equation (2) with mixing ratios of (49.97±0.18) ppbv, (0.27±0.34) ppbv and (0.03±0.30) ppbv for HONO, ClNO and NO_x, respectively (black line). The blue dashed line shows the predicted TD curve of a hypothetical mixture containing 5.0 ppbv

375 ClNO and 45.0 ppbv HONO. The insert shows a close-up of the temperature region in which TD of ClNO occurs.

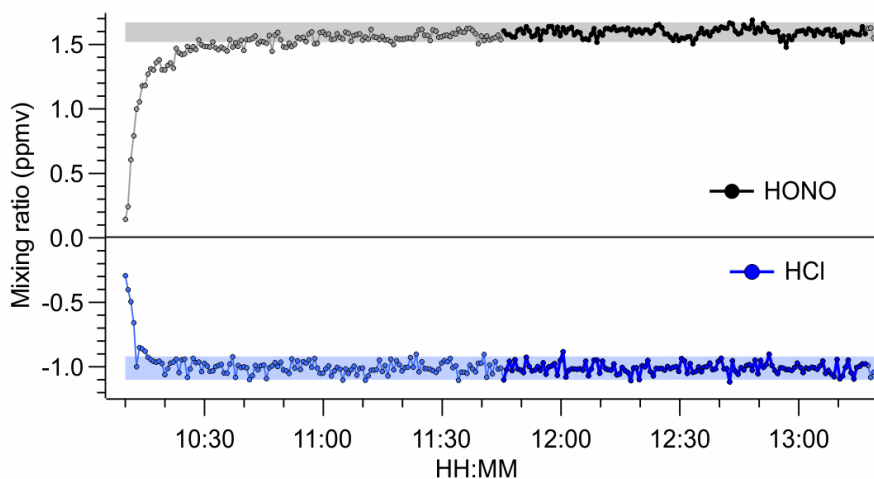



Figure 5. Time series of HONO and HCl mixing ratios derived from FTIR analysis of the undiluted HONO source output. The NaNO$_2$ was placed in line at 10:10. The shaded areas represent the average $\pm$ 2σ after the source output had stabilized (after 11:45). The 1σ precision of the HONO data was $\pm$35 ppbv and that of the HCl data was $\pm$42 ppbv. The HCl mixing ratios are an underestimate because the widths of their absorption lines are less than the FTIR's resolution of 0.5 cm^{-1}.

Table 1. Selected HONO generation systems described in the literature. n/d = not disclosed

Reference	Method	Analytical method	Result	Notes
(King and Moule, 1962; Stockwell and Calvert, 1978)	(R1)	UV absorption	mixture of NO, NO_2 and HONO	Equilibration of (R1)
(Cox, 1974)	(R1)	NO and "total NO_x" CL; HONO scrubbed w/ NaOH(0.1N)	mixture of NO, NO_2 and HONO	Equilibration of (R1)
(Cox and Derwent, 1976)	(R2)	NO and "total NO_x" CL; HONO scrubbed w/ NaOH(0.1N)	mixture of NO, NO_2 and ~50% HONO	Volatilization of "nitrous fumes" from solution of sulfuric acid and sodium nitrite
(Braman and De la Cantera, 1986)	(R3)	NO/NO_y CL analyser (heated Au tube)	mixture of NO, NO_2 and 50% - 90% HONO	Sublimation of oxalic acid on sodium nitrite
(Taira and Kanda, 1990)	(R2)	NO/NO_y CL analyser (carbon converter); HONO collected on Na_2CO_3 filters and analysed by IC	mixture containing 2% - 3% NO and 4% - 6% NO_2	Volatilization of HONO from dynamically mixed sodium nitrite and sulfuric acid solution
(Febo et al., 1995; Stutz et al., 2000)	(R4)	DOAS (NO_2, HONO); NO CL	>99.5% purity	1 L bath, 2 m reverse diffusion tube, heated stirring reactor
(Brust et al., 2000)	(R4)	FTIR (NO_2, HONO), IC (HONO)	up to 22ppm; <2% NO_2	20 cm reverse diffusion, heated reactor
(Schiller et al., 2001)	(R4)	TLDAS (HONO), FTIR (HONO) IC of KOH solutions	n/d	bubbled N_2 through 0.5 M HCl followed by reaction with solid sodium nitrite
(Perez et al., 2007)	(R4)	TD-CL, TD-LIF	observed ClNO	same setup as Febo et al. (1995)

(Perez et al., 2007)	(R2)	TD-CL, TD-LIF	>95%	flowed H_2SO_4 aerosol over $NaNO_2$ sandwiched between paper filters
(Roberts et al., 2010)	(R4)	IBBCEAS, NO_y-CL, CIMS	>95%	HCl gas cylinder (10 ppmv) and $NaNO_2$ reactor tube
(Ren et al., 2010)	(R4)	LOPAP, NO_y-CL, CIMS	n/d	Reverse diffusion tube, high [HCl] (9-12 M), tubing thickness not described
(Reed et al., 2016)	(R2)	Differential photolysis NO-CL; NO_2 CAPS; FTIR	HONO (70.4%; < 50 ppbv); NO (15%); NO_2 (12.8%); HNO_3 (1.3%)	Perm tube filled with 37% HCl placed in oven with $NaNO_2$ salt; flushed @ 1.5 slpm; <50 ppbv